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APPLICANT: Masahiro MOROOKA, Kanagawa, Japan

Yusuke SUZUKI, Kanagawa, Japan

Kazuhiro NODA, Kanagawa, Japan

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Tumminelli

TITLE: SOLID ELECTROLYTE PHOTOVOLTAIC CONVERTER AND PROCESS  
FOR PRODUCING THE SAME

Hon. Commissioner of Patents and Trademarks,  
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SIR:

CERTIFIED TRANSLATION

I, Katsuji MIYASHITA, am an official translator of the Japanese language into the English language and I hereby certify that the attached comprises an accurate translation into English of Japanese Application No. 2002-109427, filed on November 4, 2002.

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment or both, under Section 1001 of Title 18 of the United States Code and that such willful false statements may jeopardize the validity of the application or any patent issued thereon.

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Katsuji Miyashita  
Katsuji MIYASHITA

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[Inventor]

[Address] c/o Sony Corporation, 7-35, Kitashinagawa  
6-chome, Shinagawa-ku, Tokyo

[Name] Masahiro MOROOKA

[Inventor]

[Address] c/o Sony Corporation, 7-35, Kitashinagawa  
6-chome, Shinagawa-ku, Tokyo

[Name] Yusuke SUZUKI

[Inventor]

[Address] c/o Sony Corporation, 7-35, Kitashinagawa  
6-chome, Shinagawa-ku, Tokyo

[Name] Kazuhiro NODA

[Applicant for Patent]

[Id. No.] 000002185

[Name] Sony Corporation

[Agent]

[Id. No.] 100110434

[Patent Attorney]

[Name] Masaru SATO

[Application Fees]

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[Name of Document] SPECIFICATION

[Title of the Invention] SOLID ELECTROLYTE, PHOTOVOLTAIC  
DEVICE, AND MANUFACTURING METHODS THEREOF

[Claims]

[Claim 1] A solid electrolyte comprising: an electrolyte composition; and a matrix polymer,

wherein the matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[Claim 2] The solid electrolyte according to Claim 1, wherein solid electrolyte is a gel electrolyte in which a solvent is included in the electrolyte composition.

[Claim 3] The solid electrolyte according to Claim 1, wherein the solid electrolyte is an entire solid electrolyte in which no solvent is included in the electrolyte composition.

[Claim 4] The solid electrolyte according to Claim 1, wherein the solid electrolyte is a gel electrolyte in which the electrolyte composition includes an ionic liquid.

[Claim 5] The solid electrolyte according to Claim 1,

wherein the electrolyte composition includes a redox couple.

[Claim 6] The solid electrolyte according to Claim 5, wherein the redox couple is a combination of a halogen ion and a halide ion.

[Claim 7] The solid electrolyte according to Claim 6, wherein the halogen element of the redox couple is iodine.

[Claim 8] A photovoltaic device comprising, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles which carry a dye; and an electrolyte layer,

wherein the electrolyte layer includes a redox couple, an electrolyte composition; and a matrix polymer,

the matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which a solid electrolyte is to be formed.

[Claim 9] The photovoltaic device according to Claim 8, wherein the solid electrolyte is a gel electrolyte in which the electrolyte composition includes a solvent.

[Claim 10] The photovoltaic device according to Claim 8, wherein the solid electrolyte is an entire solid electrolyte

in which the electrolyte composition includes no solvent.

[Claim 11] The photovoltaic device according to Claim 8, wherein the solid electrolyte is a gel electrolyte in which the electrolyte composition includes an ionic liquid.

[Claim 12] The photovoltaic device according to Claim 8, wherein the redox couple is a combination of a halogen ion and a halide ion.

[Claim 13] The photovoltaic device according to Claim 12, wherein the halogen element of the redox couple is iodine.

[Claim 14] A method for manufacturing a photovoltaic device comprising, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles which carry a dye; and an electrolyte layer,

wherein after the photovoltaic device is assembled, a mixed solution including a first compound having at least two isocyanate groups, a second compound having at least two nucleophilic groups containing active hydrogen, and an electrolyte composition containing a redox couple is injected in the photovoltaic device, and the first compound and the second compound are solidified in the photovoltaic device by polymerization using a polyaddition reaction to form the electrolyte layer.

[Claim 15] A solid electrolyte comprising: an electrolyte composition; and a matrix polymer,

wherein the matrix polymer is formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[Claim 16] The solid electrolyte according to Claim 15, wherein solid electrolyte is a gel electrolyte in which the electrolyte composition includes a solvent.

[Claim 17] The solid electrolyte according to Claim 15, wherein the solid electrolyte is an entire solid electrolyte in which the electrolyte composition includes no solvent.

[Claim 18] The solid electrolyte according to Claim 15, wherein the solid electrolyte is a gel electrolyte in which the electrolyte composition includes an ionic liquid.

[Claim 19] The solid electrolyte according to Claim 15, wherein the electrolyte composition includes a redox couple.

[Claim 20] The solid electrolyte according to Claim 19, wherein the redox couple is a combination of a halogen ion and a halide ion.

[Claim 21] The photovoltaic device according to Claim 20, wherein the halogen element of the redox couple is iodine.

[Claim 22] A photovoltaic device comprising, between an

electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles which carry a dye; and an electrolyte layer,

wherein the electrolyte layer includes a redox couple, an electrolyte composition; and a matrix polymer,

the matrix polymer is formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which a solid electrolyte is to be formed.

[Claim 23] The photovoltaic device according to Claim 22, wherein solid electrolyte is a gel electrolyte in which the electrolyte composition includes a solvent.

[Claim 24] The photovoltaic device according to Claim 22, wherein the solid electrolyte is an entire solid electrolyte in which the electrolyte composition includes no solvent.

[Claim 25] The photovoltaic device according to Claim 22, wherein the solid electrolyte is a gel electrolyte in which the electrolyte composition includes an ionic liquid.

[Claim 26] The photovoltaic device according to Claim 22, wherein the redox couple is a combination of a halogen ion and a halide ion.

[Claim 27] The photovoltaic device according to Claim 26, wherein the halogen element of the redox couple is iodine.

[Claim 28] A method for manufacturing a photovoltaic device comprising, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles which carry a dye; and an electrolyte layer,

wherein after the photovoltaic device is assembled, a mixed solution including a first compound having at least two unsaturated double bonds, a second compound having at least two nucleophilic groups containing active hydrogen, and an electrolyte composition containing a redox couple is injected in the photovoltaic device, and the first compound and the second compound are solidified in the photovoltaic device by polymerization using a Michael addition reaction to form the electrolyte layer.

[Detailed Description of the Invention]

[0001]

[Technical Field of the Invention]

The present invention relates to solid electrolytes, and more particularly, relates to a highly reliable solid electrolyte having superior conductive properties, a photovoltaic device using this solid electrolyte, and a method of manufacturing methods thereof.

[0002]

[Description of the Related Art]

When fossil fuels, such as coal and petroleum, are used as an energy source, it has been believed that carbon dioxide which is by-produced causes global warming. In addition, when atomic energy is used, pollution may be caused by radioactive rays in some cases. In recent years in which environmental issues have been actively discussed, excessive dependence on the energies described above may result in serious problems.

[0003]

On the other hand, since solar cells, which are photovoltaic devices in which sunlight is converted into electric energy, use sunlight as an energy source, the influence of solar cells on global environment is significantly small, and hence it has been expected that solar cells become more widely used.

[0004]

Although there have been various materials which are used for forming solar cells, a large number of solar cells formed using silicon are commercially available, and the silicon solar cells can be categorized roughly into crystal silicon solar cells made of single crystal silicon or polycrystal silicon and non-crystal (amorphous) silicon solar cells. Heretofore, for the solar cells, single crystal or polycrystal silicon has been widely used.

Although the crystal silicon solar cells described above have a high conversion efficiency, which represents the performance of converting light (sunlight) energy into electric energy, as compared to that of amorphous silicon, since a large amount of energy and a long period of time are required for the crystal growth, the productivity is low, and hence the crystal silicon has been disadvantageous in view of cost.

[0005]

On the other hand, the amorphous silicon solar cells have a low conversion efficiency as compared to that of the crystal silicon solar cells; however, compared to the crystal silicon solar cells, the amorphous silicon solar cells have various advantages, such as superior light absorption properties, a wide selection range of substrate, and an easy increase in surface area. In addition, although the amorphous silicon solar cells are superior to the crystal silicon solar cells in terms of productivity, vacuum processes must be performed, and hence production facility-related loads have been still a serious problem.

[0006]

Accordingly, in order to further reduce the cost, various solar cells made of organic materials instead of silicon-bases materials have been investigated. However, when organic materials as described above are used for this

purpose, the conversion efficiency is very low, such as 1% or less, and in addition, the durability is also another problem. Among various investigations, an inexpensive solar cell using a porous semiconductor particles sensitized by a dye has been disclosed in Nature (353, p. 737-740, 1991). This solar cell is a wet type solar cell, that is, an electrochemical photocell, in which a porous thin film made of titanium oxide is used as a photoelectrode, the porous thin film being spectral-sensitized using a ruthenium complex as a sensitizing dye. The advantages of this solar cell are that an inexpensive oxide semiconductor, such as titanium oxide, can be used, light absorption of the sensitizing dye can be performed in a wide visible wavelength region of up to 800 nm, the quantum efficiency of photoelectric conversion is high, and high energy conversion efficiency can be realized. In addition, since vacuum processes are not required, a large production facility and the like are not necessary for the production.

[0007]

[Problems to be Solved by the Invention]

However, since the electrochemical photocell as described above is a wet type cell, there have been reliability problems, such as degradation in properties due to leakage and evaporation of an electrolyte solution. In order to overcome the problems described above, a gel

electrolyte composed of a polymer such as polyethylene oxide (PEO) impregnated with an electrolyte solution has also been proposed; however, since the viscosity of the electrolyte is high, and nanoscale oxide semiconductor particles are used for electrodes, the electrode pores are difficult to be filled with the electrolyte, and as a result, a problem of decrease in photoelectric conversion efficiency may arise due to low conductive properties. In addition, since cross-linking points of the above gel electrolyte form a gel by relatively weak secondary interaction such as intermolecular forces between polymers, when heat is applied thereto, a problem may also arise in that the gel is easily changed into a liquid form. Furthermore, in film formation, since a coating step and a step of removing a solvent that has a low boiling point and a low viscosity must be performed, the productivity is disadvantageously degraded.

[0008]

Hence, investigation of a chemical cross-linking type gel electrolyte has drawn a lot of attention in which a polyfunctional monomer is dissolved in an electrolyte solution and is then polymerized by applying exterior energy such as heat or active rays. The feature of this gel electrolyte is that since the viscosity of a solution containing the monomer before polymerization and a plasticizer is low, the electrode pores are easily filled

with the electrolyte solution. In addition, when a monomer solution formed by dissolving a polyfunctional monomer in an electrolyte solution is injected into a device assembled beforehand and is then in-situ gelled, a photovoltaic device can be obtained which has superior chemical bonding at an electrode interface and superior conductive properties. Furthermore, by using an ionic liquid, that is, a molten salt, instead of an electrolyte solution using a solvent, a gel electrolyte having no vapor pressure can also be realized.

[0009]

However, since iodine functions as a polymerization inhibitor in a general radical polymerization method, when an electrolyte solution contains iodine redox as an electron transport carrier, a problem may arise in that gelation cannot be in-situ performed.

[0010]

In addition, even when the iodine redox is not contained as an electron transport carrier, degradation of a sensitizing dye and an electrolyte layer inevitably occurs by application of heat and/or active rays in polymerization, and hence degradation of the photoelectric conversion properties disadvantageously occurs.

[0011]

Hence, the present invention was made in consideration

of the conventional situation described above, and an object of the present invention is to provide a highly reliable solid electrolyte having superior conductive properties, a photovoltaic device using the above solid electrolyte, and a method of manufacturing methods thereof.

[0012]

[Means for Solving the Problems]

Accordingly, the present invention was made to solve the problems described above by a novel polymerization method without using a radical polymerization method.

[0013]

That is, a solid electrolyte of the present invention which achieves the above object is a solid electrolyte comprising an electrolyte composition and a matrix polymer, the matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0014]

In addition, another solid electrolyte of the present invention which achieves the above object is a solid electrolyte comprising an electrolyte composition and a

matrix polymer, the matrix polymer is formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0015]

The solid electrolytes of the present invention described above are each preferably a gel electrolyte in which a solvent is included in the electrolyte composition.

[0016]

In addition, the solid electrolytes of the present invention described above are each preferably an entire solid electrolyte in which no solvent is included in the electrolyte composition.

[0017]

In addition, the solid electrolytes of the present invention described above are each preferably a gel electrolyte in which the electrolyte composition includes an ionic liquid.

[0018]

In addition, in each of the solid electrolytes of the present invention described above, the electrolyte

composition preferably includes a redox couple, that is, the above solid electrolyte is preferably a solid electrolyte for photovoltaic devices.

[0019]

In addition, when the solid electrolyte of the present invention described above is a solid electrolyte for photovoltaic devices, the solid electrolyte is preferably used for a photovoltaic device having a redox couple formed in combination of a halogen ion and a halide ion, and in the case described above, in particular, the solid electrolyte is preferably used for a photovoltaic device in which the halogen element is iodine.

[0020]

A photovoltaic device of the present invention, which achieves the above object, is a photovoltaic device comprising, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles carrying a dye and an electrolyte layer. The electrolyte layer includes a redox couple, an electrolyte composition, and a matrix polymer, and the matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen. In addition, the polymerization is

performed after a precursor for the matrix polymer is brought into contact with a surface on which a solid electrolyte is to be formed.

[0021]

In addition, another photovoltaic device of the present invention, which achieves the above object, is a photovoltaic device comprising, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles carrying a dye and an electrolyte layer. The electrolyte layer includes a redox couple, an electrolyte composition, and a matrix polymer, and the matrix polymer is formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen. In addition, the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which a solid electrolyte is to be formed.

[0022]

In each of the photovoltaic devices of the present invention described above, the solid electrolyte is preferably a gel electrolyte in which a solvent is included in the electrolyte composition.

[0023]

In addition, in each of the photovoltaic devices of the present invention described above, the solid electrolyte is preferably an entire solid electrolyte in which no solvent is included in the electrolyte composition.

[0024]

In addition, in each of the photovoltaic devices of the present invention described above, the solid electrolyte is preferably a gel electrolyte in which the electrolyte composition includes an ionic liquid.

[0025]

In addition, in each of the photovoltaic devices of the present invention described above, the electrolyte composition of the solid electrolyte preferably includes a redox couple, that is, the solid electrolytes are each preferably a solid electrolyte for photovoltaic devices.

[0026]

In addition, the above photovoltaic devices of the present invention are each preferably used when the redox couple described above is a combination of a halogen ion and a halide ion, and in this case, as the halogen element, iodine is particularly preferable.

[0027]

A method for manufacturing a photovoltaic device according to the present invention, which achieves the above object, is a method for manufacturing a photovoltaic device

which has, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles carrying a dye and an electrolyte layer. In the method described above, after the photovoltaic device is assembled, a mixed solution including a first compound having at least two isocyanate groups, a second compound having at least two nucleophilic groups containing active hydrogen, and an electrolyte composition containing a redox couple is injected in the photovoltaic device, and the first compound and the second compound are solidified in the photovoltaic device by polymerization using a polyaddition reaction to form the electrolyte layer.

[0028]

In addition, another method for manufacturing a photovoltaic device according to the present invention, which achieves the above object, is a method for manufacturing a photovoltaic device which has, between an electrode formed on a surface of a transparent substrate and a counter electrode, a semiconductor layer composed of semiconductor particles carrying a dye and an electrolyte layer. In the method described above, after the photovoltaic device is assembled, a mixed solution including a first compound having at least two unsaturated double bonds, a second compound having at least two nucleophilic

groups containing active hydrogen, and an electrolyte composition containing a redox couple is injected in the photovoltaic device, and the first compound and the solid compound are solidified in the photovoltaic device by polymerization using a Michael addition reaction to form the electrolyte layer.

[0029]

According to the methods for manufacturing a photovoltaic device, of the present invention, the photovoltaic devices described above can be efficiently and reliably formed.

[0030]

[Description of the Embodiments]

Hereinafter, a solid electrolyte of the present invention, a photovoltaic device using this solid electrolyte, and manufacturing methods thereof will be described in detail. However, the present invention is not limited to the following descriptions, and modification may be optionally made without departing from the spirit and the scope of the present invention.

[0031]

The solid electrolyte of the present invention is a solid electrolyte having an electrolyte composition and a matrix polymer. In this solid electrolyte, the matrix polymer is formed by polymerization using a polyaddition

reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0032]

Fig. 1 is a cross-sectional view showing the structure of a dye-sensitized solar cell which is a photovoltaic device formed in accordance with the present invention described above. As shown in Fig. 1, a dye-sensitized solar cell 1 is formed of a transparent substrate 2, a transparent electrode 3, a semiconductor layer 4, a solid electrolyte 5, a platinum layer 6 processed by platinum-chloride treatment, a transparent electrode 7, and a transparent substrate 8.

[0033]

Materials for the transparent substrates 2 and 8 are not particularly limited as long as having transparency, and for example, a glass substrate may be used.

[0034]

The transparent electrodes 3 and 7 are electrodes formed using a transparent material and provided under the transparent substrate 2. As materials for the electrodes, any materials may be used as long as having conductivity and transparency; however, since simultaneously having superior

conductivity, transparency, and heat resistance, tin-based oxides and the like are preferably used, and in addition, ITO is preferable in view of cost. In addition, the transparent electrode 7 is not always necessary to be provided and may be formed whenever necessary.

[0035]

The semiconductor layer 4 is formed on the transparent electrode 3 by sintering semiconductor particles carrying a dye, and the dye absorbs light which passes through the transparent substrate 2 and the transparent electrode 3 and is then incident on the semiconductor layer 4. The semiconductor particles are semiconductor fine particles adsorbing a dye, and as the semiconductor fine particles, besides elemental semiconductors represented by silicon, compound semiconductors, compounds having a perovskite structure, and the like may also be used. As the semiconductors described above, an n-type semiconductor is preferable in which conduction band electrons serve as carriers under photoexcitation conditions to convey an anode current. In particular, for example, there may be mentioned  $\text{TiO}_2$  (titania),  $\text{SnO}_2$ ,  $\text{ZnO}$ ,  $\text{WO}_3$ ,  $\text{Nb}_2\text{O}_5$ , and  $\text{TiSrO}_3$ , and anatase  $\text{TiO}_2$  is particularly preferable. In addition, the types of semiconductors are not limited to those mentioned above, and the aforementioned materials may be used alone or in combination.

[0036]

A film-forming method of the semiconductor layer 4 is not particularly limited; however, in consideration of properties, handiness, manufacturing cost, and the like, a wet film-forming method is preferably used for semiconductor fine particles, and a method is preferably used having the steps of preparing a paste in which semiconductor fine particles in a powder form or a sol form are uniformly dispersed in a solvent such as water, and applying the paste onto a substrate provided with a transparent conductive film thereon. An application method is not particularly limited, and application may be performed using various known methods. As the known methods, for example, there may be mentioned dipping, spraying, wire-bar coating, spin coating, roller coating, blade coating, and gravure coating, and in addition, as wet printing methods, for example, relief printing, off-set printing, gravure printing, intaglio printing, rubber plate printing, and screen printing may be mentioned.

[0037]

As the crystal structure of crystal titanium oxide, an anatase structure is preferable in view of photocatalytic activity. As the anatase titanium oxide, a commercially available powder, sol, and slurry may be used, and alternatively, a titanium oxide having a predetermined particle size may be formed by a known method in which, for

example, a titanium oxide alkoxide is hydrolyzed. When a commercially available powder is used, secondary aggregation of the particles is preferably resolved, and hence the particles are preferably pulverized by the use of a mortar, a ball mill, or the like when a coating solution is prepared. In this step, in order to prevent re-aggregation of the particles which were processed once for resolving the secondary aggregation, acetyl acetone, hydrochloric acid, nitric acid, a surfactant, a chelating agent, or the like may be added. In addition, in order to increase viscosity, various thickening agents, such as a polymer including polyethylene oxide or polyvinyl alcohol, and a cellulose-based thickening agent, may be added.

[0038]

Although the particle diameter of the semiconductor fine particles is not particularly limited, the average particle diameter of the primary particles is preferably in the range of 1 to 200 nm and particularly preferably in the range of 5 to 100 nm. In addition, at least two types of particles having a size larger than that of the above semiconductor fine particles may be mixed together for scattering incident light, so that the quantum yield can be improved. In this case, the average particle size of the particles which are additionally mixed is preferably in the range of 20 to 500 nm.

[0039]

The semiconductor layer preferably has a large surface area so that a large amount of dyes is adsorbed thereon. Hence, the surface area obtained when the semiconductor fine particles are applied onto a supporting body is preferably 10 times or more the projection area thereof and is more preferably 100 times or more. The upper limit of this surface area is not particularly specified; however, in general, the surface area is approximately 1,000 times the projection area. In general, when the thickness of a layer containing semiconductor fine particles is increased, the light-capturing rate is increased since the dye amount carried per unit projection area is increased; however, since the diffusion length of injected electrons is also increased, a loss caused by charge recombination is also increased. Hence, in the semiconductor layer, a preferable thickness is present; however, the thickness is generally in the range of 0.1 to 100  $\mu\text{m}$ , is more preferably in the range of 1 to 50  $\mu\text{m}$ , and is particularly preferably in the range of 3 to 30  $\mu\text{m}$ .

[0040]

After the semiconductor fine particles are applied onto the supporting body, in order to electronically bring the fine particles into contact with each other and to improve the film strength and the adhesion with the substrate,

firing is preferably performed. The firing temperature range is not particularly limited; however, when the temperature is excessively increased, the resistance of the substrate may be increased so that it is melted in some cases, and hence the temperature is generally in the range of 40 to 700°C and is more preferably in the range of 40 to 650°C. In addition, the firing time is also not particularly limited; however, it is generally in the range of approximately 10 minutes to 10 hours. After the firing, in order to increase the surface area of the semiconductor particles, to remove impurities of the semiconductor layer, and to improve the electron injection efficiency from the dye into the semiconductor particles, for example, chemical plating using an aqueous solution containing tetrachloride titanium or electrochemical plating treatment using an aqueous solution containing trichloride titanium may be performed.

[0041]

As the dye adsorbed onto the semiconductor fine particles, for example, a ruthenium-based dye is preferably used. However, the dye adsorbed onto the semiconductor fine particles is not particularly limited as long as having a charge separation function and a sensitizing effect. Hence, besides the ruthenium-based dye, for example, there may be mentioned xanthene dyes such as rhodamine B, rose bengal,

eosine, and erythrosine; cyanine dyes such as quinocyanine and cryptocyanine; basic dyes such as phenosafranine, capri blue, thionine, and methylene blue; porphyrin compounds such as chlorophyll, zinc porphyrin, and magnesium porphyrin; complex compounds such as azo dyes, phthalocyanine compounds, and Ru trisbipyridyl; anthraquinone dyes; and polycyclic quinone dyes. The dyes mentioned above may be used alone or in combination.

[0042]

A method for adsorbing the above dye onto the semiconductor layer is not particularly limited; however, after the dye is dissolved in a solvent, such as alcohol, nitrile, nitromethane, halogenated hydrocarbon, ether, dimethyl sulfoxide, amide, N-methyl pyrrolidone, 1,3-dimethyl imidazolidinone, 3-methyl oxazolidinone, ester, carbonic acid ester, ketone, hydrocarbon, or water, an electrode having a semiconductor layer may be immersed therein, or a dye solution may be applied onto a semiconductor layer. In addition, in order to reduce the association between dye elements, deoxycholic acid, chenodeoxycholic acid, or the like may be added. In addition, a UV absorber may also be used.

[0043]

In addition, in order to facilitate removal of an excessively adsorbed dye, after the dye is adsorbed, the

surfaces of the semiconductor fine particles may be processed using an amine. As the amine, for example, there may be mentioned pyridine, 4-tert-butylpyridine, or polyvinylpyridine. When being in the liquid form, the amine mentioned above may be used without being processed or may be used after being dissolved in an organic solvent.

[0044]

The solid electrolyte 5 is formed of a gel electrolyte or an entire solid electrolyte, which functions as a carrier transport layer, and is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen to form a polymer material having a cross-linking matrix. Accordingly, in this dye-sensitized solar cell, when the solid electrolyte 5 formed of a gel electrolyte or an entire solid electrolyte is used for the electrolyte layer as described above, a decrease of the electrolyte layer due to leakage and evaporation of liquid, which occurs when an electrolyte solution is used, can be prevented, and hence a dye-sensitized solar cell having superior cell properties and reliability can be realized.

[0045]

One of the first compound and the second compound preferably has ether, ester, carbonate, alkyl,

perfluorocarbon, nitrile, tertiary amine, or the like at the main and the side chains. The number of the isocyanate groups of the first compound is preferably at least two; however, when the number of the nucleophilic groups containing active hydrogen of the second compound is two, the number of the isocyanate groups of the first compound must be at least three. As is the case described above, the number of the nucleophilic groups containing active hydrogen of the second compound is preferably at least two; however, when the number of the isocyanate groups of the first compound is two, the number of the nucleophilic groups containing active hydrogen of the second compound must be at least three.

[0046]

In addition, compounds having different structures may be used alone or in combination. As the first compound, for example, tolylene 2,4-diisocyanate, 4,4'-diphenylmethane diisocyanate, hexamethylene diisocyanate, isophorone diisocyanate, hydrogenated 4,4'-diphenylmethane diisocyanate, a trimer of hexamethylene diisocyanate, or a polymer of isocyanate ethylmethacrylate may be mentioned by way of example. In order to improve light resistance, as the first compound, aliphatic isocyanate compounds are preferably selected and may be used alone or in combination.

[0047]

In addition, as the second compound, for example, there may be mentioned a polyol compound, such as diol, triol, or tetraol; a polyamine compound such as diamine, triamine, or tetramine as is the case described above; or a polycarboxylic acid such as dicarboxylic acid, tricarboxylic acid, or teracarboxylic acid. The compounds mentioned above may be used alone or in combination.

[0048]

In addition, in order to efficiently carry out the polyaddition reaction, a catalyst may be used. As the catalyst, in general, a known catalyst used for polyurethane synthesis may be mentioned, such as a tin-based catalyst including dibutyl tin dilaurate or an amine-based catalyst; however, the catalyst is not specifically limited to those mentioned above, and the catalysts mentioned above may be used alone or in combination. In addition, when the catalyst is used, the addition amount thereof is 1 percent by weight or less and is preferably 0.1 percent by weight or less.

[0049]

When the solid electrolyte 5 is a gel electrolyte, the gel electrolyte is composed of an electrolyte composition containing a solvent and the above cross-linking matrix, and the ratio of the cross-linking matrix is 3 to 50 percent by weight of the gel electrolyte. As the solvent forming the

electrolyte composition, for example, there may be mentioned water, alcohol, ether, ester, carbonic acid ester, lactone, carboxylic acid ester, phosphate triester, heterocyclic compound, nitrile, ketone, amide, nitormethane, halogenated hydrocarbon, dimethyl sulfoxide, sulfolane, N-methyl pyrrolidone, 1,3-dimethyl imidazolidinone, 3-methyl oxazolidinone, or hydrocarbon. However, the solvent is not limited to those mentioned above, and the aforementioned solvents may be used alone or in combination. In addition, among the solvents mentioned above, aprotic nonaqueous solvents are more preferable.

[0050]

When the ratio of the electrolyte composition to the gel electrolyte is increased, the ion conductivity is increased, and on the other hand, the mechanical strength is decreased. On the contrary, as the ratio of the electrolyte composition to the gel electrolyte 5 is decreased, the mechanical strength is increased, but the ion conductivity is decreased. Hence, the ratio of the electrolyte composition is preferably 50 to 97 percent by weight of the gel electrolyte and is more preferably 80 to 95 percent by weight thereof.

[0051]

As the electrolytes used for the gel electrolyte, in addition to the combination of  $I_2$  and a metal iodide or an

organic iodide and the combination of  $\text{Br}_2$  and a metal bromide or an organic bromide, for example, metal complexes such as ferrocyanate/ferricyanate and ferrocene/ferricinium ion; sulfur compounds, such as sodium polysulfide and alkylthiol/alkyldisulfide; viologne dyes; and hydroquinone/quinone may be used.

[0052]

As cations of the above metal compounds, Li, Na, K, Mg, Ca, Cs, and the like are preferable, and as cations of the above organic compounds, quaternary ammonium compounds, such as tetraalkylammonium, pyridinium, and imidazolium, are preferable. However, the cations are not limited to those mentioned above, and the above cations may be used alone or in combination. Among those mentioned above, the combination of  $\text{I}_2$  and LiI or a quaternary ammonium compound such as imidazolium iodide is preferably used as the electrolyte. The concentration of the electrolyte salt is preferably 0.05 to 5 M to the solvent and is more preferably 0.2 to 1 M. The concentration of  $\text{I}_2$  or  $\text{Br}_2$  is preferably 0.0005 to 1 M and is more preferably 0.001 to 0.1 M. In addition, in order to improve the open-circuit voltage and short-circuit current, various additives, such as 4-tert-butylpyridine and a carboxylic acid, may also be added.

[0053]

In addition, when the solid electrolyte 5 is an entire

solid electrolyte, the entire solid electrolyte is formed, for example, of the cross-linking matrix containing iodine redox. As the electrolytes, in addition to the combination of  $I_2$  and a metal iodide or an organic iodide and the combination of  $Br_2$  and a metal bromide or an organic bromide, for example, metal complexes such as ferrocyanate/ferricyanate and ferrocene/ferricinium ion; sulfur compounds, such as sodium polysulfide and alkylthiol/alkyldisulfide; viologne dyes; and hydroquinone/quinone may be used.

[0054]

As cations of the above metal compounds, Li, Na, K, Mg, Ca, Cs, and the like are preferable, and as cations of the above organic compounds, quaternary ammonium compounds, such as tetraalkylammonium, pyridinium, and imidazolium, are preferable. However, the cations are not limited to those mentioned above, and the above cations may be used alone or in combination. Among those mentioned above, an electrolyte formed in combination of  $I_2$  and LiI or a quaternary ammonium compound, such as imidazolium iodide, is preferable. The concentration of the electrolyte salt is preferably 0.05 to 5 M to the monomer and is more preferably 0.2 to 1 M. In addition, the concentration of  $I_2$  or  $Br_2$  is preferably 0.0005 to 1 M and is more preferably 0.001 to 0.1 M.

[0055]

Since the solid electrolyte 5 having the structure as described above is formed of a gel electrolyte or an entire solid electrolyte, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a highly reliable electrolyte is obtained.

[0056]

In addition, since being a polymer formed by the polyaddition reaction between the above first and second compounds and having a cross-linking structure, the matrix polymer of the solid electrolyte 5 is not changed into a liquid form by heating, and hence superior mechanical properties and durability can be obtained.

[0057]

In addition, since this solid electrolyte 5 is formed by polymerization after a precursor therefor, which is in a state before polymerization and which has fluidity, is brought into contact with a surface on which the solid electrolyte is to be formed, that is, with an electrode surface, the pores in the electrode surface are sufficiently filled with the electrolyte, and the chemical bonding condition at the electrochemical interface between the solid electrolyte and the electrode surface is improved, thereby obtaining superior conductive properties.

[0058]

In addition, since this solid electrolyte 5 is formed by polymerization using the polyaddition reaction between the above first and second compounds, the use of heat and/or active rays is not necessary in polymerization. Accordingly, the electrolyte composition is not degraded by heat and/or active rays in production, and hence an electrolyte having superior conductive properties can be realized. In addition, the manufacturing process is simple, and the productivity is superior.

[0059]

In addition, since this solid electrolyte is formed by the polyaddition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte can be easily formed, and in addition, since being in-situ formed in a cell device, the solid electrolyte is preferably applied to a photovoltaic device or the like which uses an iodine redox couple.

[0060]

The platinum layer 6 is a counter electrode formed of any optional material as long as it has conductive properties. In addition, even in the case in which an insulating material is used, when having a conductive layer at a side facing a semiconductor electrode, the insulating

material may also be used. However, an electrochemically stable material is preferably used, and in particular, platinum, gold, carbon, or the like is preferably used. In addition, in order to improve a catalytic effect of oxidation-reduction, the surface at a side facing the semiconductor electrode preferably has a fine structure and an increased surface area. For example, in the case of platinum, platinum in a platinum black state is preferably used, and in the case of carbon, carbon having a porous structure is preferably used. The platinum in a platinum black state may be obtained by an anodization method or treatment using a chloroplatinic acid solution, and the carbon having a porous structure may be formed by sintering carbon fine particles or by firing an organic polymer.

[0061]

The dye-sensitized solar cell 1 having the structure as described above is operated as described below. Light incident from the transparent electrode 2 at a transparent electrode 3 side excites the dye carried on the surface of the semiconductor layer 4, and the dye smoothly transports electrons to the semiconductor fine particles of the semiconductor layer 4. In addition, the dye which lost electrons receives electrons from ions of the gel electrolyte 5. Molecules which transported electrons again receive electrons at the platinum layer 6 which is the

counter electrode.

[0062]

In addition, the above dye-sensitized solar cell 1 can be formed as described below. Hereinafter, the case in which a gel electrolyte is used as the solid electrolyte 5 will be described by way of example.

[0063]

First, on one major surface of the transparent substrate 2, the transparent electrode 3 is formed, and on the transparent electrode 3, the semiconductor layer 4 carrying the dye is formed, thereby forming the semiconductor electrode. In addition, on one major surface of the transparent substrate 8, the transparent electrode 7 is formed, and on the transparent electrode 7, the platinum layer 6 processed by platinum-chloride treatment is formed. Subsequently, after the semiconductor layer 4 carrying the dye and the platinum layer 6 are disposed to face each other, side surfaces of the transparent substrate 2 and the transparent substrate 7 are sealed so that the semiconductor electrode and the counter electrode are not brought into contact with each other. In this step, the distance between the semiconductor electrode and the counter electrode is not particularly limited; however, the distance is generally 1 to 100  $\mu\text{m}$  and is more preferably 1 to 30  $\mu\text{m}$ . When this distance is excessively large, the conductivity is decreased,

and as a result, the amount of photocurrent is unfavorably decreased.

[0064]

In addition, although the sealing method is also not specifically limited, a material having light resistance, insulating properties, and humidity resistance is preferably used, and for example, an epoxy resin, a UV curable resin, an acrylic adhesive, ethylene vinyl acetate (EVA), a ceramic, or a heat sealable film may be used. In addition, the photovoltaic device of the present invention must have an inlet opening through which a solution before gelation is injected. The place of the inlet opening is not particularly limited as long as it is not on the semiconductor layer 4 carrying the dye and part of the counter electrode that faces the semiconductor layer 4, and hence the inlet opening may be provided at an optional place.

[0065]

Next, a mixed solution is prepared which serves as the precursor for the gel electrolyte 5. First, the second compound having at least two nucleophilic groups containing active hydrogen is dissolved in an electrolyte solution which is the electrolyte composition. Subsequently, the first compound having at least two isocyanate groups is dissolved, thereby preparing the mixed solution before gelation. In this process, the order of the dissolution of

the compounds is not particularly limited; however the isocyanate groups are liable to perform an addition reaction with iodine in the electrolyte composition, the compounds are preferably dissolved in that order described above.

[0066]

The mixed solution is injected into the dye-sensitized solar cell 1 assembled as described above. The method for injecting the mixed solution is not specifically limited; however, a method is preferably performed in which after the compounds are thoroughly dissolved, the mixed solution is injected inside the dye-sensitized solar cell 1 which is sealed beforehand and which is provided with the inlet opening for the mixed solution. In this case, a method in which after several droplets of the mixed solution are dripped in the inlet opening, the mixed solution is injected using a capillary phenomenon can be easily performed. In addition, whenever necessary, the injection of the mixed solution may be performed at a reduced pressure. After the mixed solution is entirely injected, a mixed solution before gellation remaining at the inlet opening is removed, and this inlet opening is then sealed. This sealing method is also not particularly limited, and whenever necessary, sealing may be performed by bonding a glass plate with a sealing agent.

[0067]

After injected into the dye-sensitized solar cell 1, the mixed solution before gelation is preferably allowed to stand still until the polymerization between the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen is completed by the polyaddition reaction. The standing time is not particularly limited; however, the time in which the mixed solution injected inside the dye-sensitized solar cell 1 completely loses its viscosity and the gelation is completed is in the range of approximately 1 minute to 48 hours. However, the time described above may vary in some cases in accordance with various conditions, that is, in accordance with the selections of the first and the second compounds, the electrolyte, the solvent, and the like. In addition, the ambient temperature for the standing is not particularly limited; however, it is generally in the range of 0 to 120°C and is preferably in the range of 0 to 80°C for reducing the influence on the dye and the electrolyte.

[0068]

In addition, in the dye-sensitized solar cell 1, in order to increase power generation efficiency, the surface of the transparent substrate 2 at the side on which light is incident may be processed by antireflection (AR) treatment. In addition, by processing the surface of the dye-sensitized

solar cell 1 by grooving, texturing, or the like, the efficiency of the use of incident light may be increased. In addition, in order to effectively use light passing through the dye-sensitized solar cell 1, the bottom layer of the counter electrode may be provided with a metal, such as Al or Ag, or an alloy thereof having a high reflectance by sputtering, deposition, or the like.

[0069]

As described above, the dye-sensitized solar cell 1 can be formed.

[0070]

In the dye-sensitized solar cell 1 thus formed, since the solid electrolyte 5 described above is provided, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a dye-sensitized solar cell having superior reliability can be realized.

[0071]

In addition, in this dye-sensitized solar cell 1, since the matrix polymer of the solid electrolyte 5 is a polymer formed of the above first and second compounds by the polyaddition reaction and has a chemically cross-linking structure, the solid electrolyte 5 is not changed into a liquid form by heating, and hence a dye-sensitized solar cell having superior mechanical properties and durability

can be realized.

[0072]

In addition, in the case in which a solid electrolyte is used for a dye-sensitized solar cell, when the solid electrolyte is formed beforehand and is then provided so as to be in close contact with a semiconductor layer for forming the dye-sensitized solar cell, since the semiconductor layer and the solid electrolyte each having its own shape, that is, each having a solidified surface shape, are brought into contact with each other, adhesion between the semiconductor layer and the solid electrolyte is not satisfactory. As a result, due to the insufficient contact between the semiconductor layer and the solid electrolyte, a problem may arise in that the photoelectric conversion efficiency is decreased.

[0073]

However, in this dye-sensitized solar cell 1, since the solid electrolyte 5 is formed by polymerization after the precursor for the solid electrolyte 5, which is in a state before polymerization and which has fluidity, is injected into the dye-sensitized solar cell, the pores in the electrode surface are sufficiently filled with the electrolyte, and in addition, the adhesion of the solid electrolyte 5 with the semiconductor fine particles of the semiconductor layer 4 and the platinum layer 6 used as the

counter electrode can be increased. Accordingly, since the contact between the semiconductor layer 4 and the solid electrolyte 5 can be satisfactorily secured, the chemical bonding condition at the electrochemical interface between the solid electrolyte 5 and the electrode surface is improved, thereby realizing a dye-sensitized solar cell having superior photoelectric conversion properties.

[0074]

In addition, in this dye-sensitized solar cell 1, since the solid electrolyte 5 is formed by polymerization using the polyaddition reaction, the use of heat and/or active rays is not necessary. Accordingly, the electrolyte composition is not degraded by the use of heat and/or active rays in production of the solid electrolyte, and hence a photovoltaic device having superior photoelectric conversion properties can be realized. In addition, the manufacturing process is simple, and the productivity is superior.

[0075]

In addition, in this dye-sensitized solar cell 1, since this solid electrolyte 5 is formed by the polyaddition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte 5 can be easily formed, and in addition, the

solid electrolyte 5 can be in-situ formed in a cell device. Hence, according to this dye-sensitized solar cell 1, a dye-sensitized solar cell having superior photoelectric conversion properties can be easily and reliably formed.

[0076]

In the solid electrolyte 5 described above, by using an ionic liquid, that is, a molten salt, instead of the electrolyte composition containing a solvent, a gel electrolyte having no vapor pressure can also be formed.

[0077]

Next, another solid electrolyte and photovoltaic device, according to the present invention, and manufacturing methods thereof will be described in detail. Fig. 2 is a cross-sectional view showing the structure of another dye-sensitized solar cell formed in accordance with the present invention. As shown in Fig. 2, a dye-sensitized solar cell 11 is formed of the transparent substrate 2, the transparent electrode 3, the semiconductor layer 4, a solid electrolyte 15, the platinum layer 6 processed by platinum-chloride treatment, the transparent electrode 7, and the transparent substrate 8. In this dye-sensitized solar cell 11, the same reference numerals of the constituent elements of the above dye-sensitized solar cell 1 shown in Fig. 1 are used to designate the same constituent elements, and descriptions thereof in detail will be omitted. A difference between the

dye-sensitized solar cell 11 and the dye-sensitized solar cell 1, that is, in more particular, the solid electrolyte 15, will only be described.

[0078]

The solid electrolyte 15 is formed of a gel electrolyte which functions as a carrier transport layer and is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen to form a polymer material having a cross-linking matrix. In addition, the solid electrolyte 15 is formed of an ionic liquid containing a redox couple and 3 to 50 percent by weight of the cross-linking matrix described above. Accordingly, in this dye-sensitized solar cell, when the solid electrolyte 15 formed of the gel electrolyte is used as an electrolyte layer, the decrease of the electrolyte layer due to leakage and evaporation of liquid, which occurs when an electrolyte solution is used, can be prevented, and hence a dye-sensitized solar cell having superior cell properties and reliability can be realized.

[0079]

In addition, one of the first compound and the second compound preferably has ether, ester, carbonate, alkyl, perfluorocarbon, nitrile, tertiary amine, or the like at the

main and the side chains. The number of the isocyanate groups of the first compound is preferably at least two; however, when the number of the nucleophilic groups containing active hydrogen of the second compound is two, the number of the isocyanate groups of the first compound must be at least three. As is the case described above, the number of the nucleophilic groups containing active hydrogen of the second compound is preferably at least two; however, when the number of the isocyanate groups of the first compound is two, the number of the nucleophilic groups containing active hydrogen of the second compound must be at least three.

[0080]

In addition, compounds having different structures may be used alone or in combination. As the first compound, for example, tolylene 2,4-diisocyanate, 4,4'-diphenylmethane diisocyanate, hexamethylene diisocyanate, isophorone diisocyanate, hydrogenated 4,4'-diphenylmethane diisocyanate, a trimer of hexamethylene diisocyanate, or a polymer of isocyanate ethylmethacrylate may be mentioned by way of example. In order to improve light resistance, as the first compound, aliphatic isocyanate compounds are preferably selected and may be used alone or in combination.

[0081]

In addition, as the second compound, for example, there

may be mentioned a polyol compound, such as diol, triol, or tetraol; a polyamine compound such as diamine, triamine, or tetramine as is the case described above; or a polycarboxylic acid such as dicarboxylic acid, tricarboxylic acid, or teracarboxylic acid. The compounds mentioned above may be used alone or in combination.

[0082]

In addition, in order to efficiently carry out the polyaddition reaction, a catalyst may be used. As the catalyst, in general, a known catalyst used for polyurethane synthesis may be mentioned, such as a tin-based catalyst including dibutyl tin dilaurate or an amine-based catalyst; however, the catalyst is not specifically limited to those mentioned above, and the catalysts mentioned above may be used alone or in combination. In addition, when the catalyst is used, the addition amount thereof is 1 percent by weight or less and is preferably 0.1 percent by weight or less.

[0083]

As the ionic liquid, for example, there may be mentioned a pyridinium salt, an imidazolium salt, or a triazorium salt; however, the ionic liquid is not limited thereto. In addition, those mentioned above may be used alone or in combination. The melting point of the ionic liquid is preferably 100°C or less, more preferably 80°C or

less, and particularly preferably 60°C or less.

[0084]

When the ratio of the ionic liquid to the gel electrolyte is increased, the ion conductivity is increased thereby, but the mechanical strength is decreased. On the other hand, as the ratio of the ionic liquid to the gel electrolyte is decreased, the mechanical strength is increased; however, the ion conductivity is decreased. Hence, the ratio of the ionic liquid is preferably 50 to 97 percent by weight of the gel electrolyte and more preferably 80 to 95 percent by weight thereof.

[0085]

As the redox couple used for the gel electrolyte of the present invention, the combination of I<sub>2</sub> and an ionic liquid of iodide or the combination of Br<sub>2</sub> and an ionic liquid of bromide is preferable, and in particular, the combination of I<sub>2</sub> and the ionic liquid is preferable. The concentration of the redox species to the total electrolyte is preferably 0.1 to 20 percent by weight and is more preferably 0.2 to 5 percent by weight.

[0086]

In addition, besides the materials mentioned above, as an counter anion of the ionic liquid, for example, there may be mentioned a halogen ion, such as Cl<sup>-</sup> or Br<sup>-</sup>, NSC<sup>-</sup>, BF<sub>4</sub><sup>-</sup>, PF<sub>6</sub><sup>-</sup>, ClO<sub>4</sub><sup>-</sup>, (CF<sub>3</sub>SO<sub>2</sub>)<sub>2</sub>N<sup>-</sup>, (CF<sub>3</sub>CF<sub>2</sub>SO<sub>2</sub>)<sub>2</sub>N<sup>-</sup>, CF<sub>3</sub>SO<sub>3</sub><sup>-</sup>, CF<sub>3</sub>COO<sup>-</sup>,

$\text{Ph}_4\text{B}^-$ ,  $(\text{CF}_3\text{SO}_2)_3\text{C}^-$ , or  $\text{F}(\text{HF})_n$ , and among those mentioned above,  $(\text{CF}_3\text{SO}_2)_2\text{N}^-$  or  $\text{BF}_4^-$  is preferable.

[0087]

In addition, as the redox couples in this case, besides the combination with a metal iodide or an organic iodide and the combination of  $\text{Br}_2$  and a metal bromide or an organic bromide, for example, metal complexes such as ferrocyanate/ferricyanate and ferrocene/ferricinium ion; sulfur compounds, such as sodium polysulfide and alkylthiol-alkyldisulfide; viologne dyes; and hydroquinone/quinone may be used.

[0088]

As cations of the above metal compounds, Li, Na, K, Mg, Ca, Cs, and the like are preferable, and as cations of the above organic compounds, quaternary ammonium compounds, such as tetraalkylammonium, pyridinium, and imidazolium, are preferable. However, the cations are not limited to those mentioned above, and the above cations may be used alone or in combination. Among those mentioned above, an electrolyte formed in combination of  $\text{I}_2$  and a quaternary ammonium compound, such as imidazolium iodide, is preferable.

[0089]

The concentration of the electrolyte salt is preferably 0.05 to 5 M with respect to the ionic liquid and is more preferably 0.2 to 1 M. The concentration of  $\text{I}_2$  or  $\text{Br}_2$  is

preferably 0.0005 to 1 M and is more preferably 0.001 to 0.1 M. In addition, in order to improve the open-circuit voltage and short-circuit current, various additives, such as 4-tert-butylpyridine and a carboxylic acid, may also be added.

[0090]

Since the solid electrolyte 15 having the structure as described above is formed of a gel electrolyte, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a highly reliable electrolyte is obtained.

[0091]

In addition, since being a polymer formed by the polyaddition reaction between the above first and the second compounds and having a chemically cross-linking structure, the matrix polymer of the solid electrolyte 15 is not changed into a liquid form by heating, and hence superior mechanical properties and durability can be obtained.

[0092]

In addition, since the solid electrolyte 15 is formed by polymerization after a precursor therefor, which is in a state before polymerization and which has fluidity, is brought into contact with a surface on which the solid electrolyte is to be formed, that is, with an electrode surface, the pores in the electrode surface are sufficiently

filled with the electrolyte, and in addition, the chemical bonding condition at the electrochemical interface between the solid electrolyte and the electrode surface is improved, thereby obtaining superior conductive properties.

[0093]

In addition, since the solid electrolyte 15 is formed by polymerization using the polyaddition reaction between the above first and second compounds, the use of heat and/or active rays is not necessary in polymerization. Accordingly, the electrolyte composition is not degraded by heat and/or active rays in production, and hence an electrolyte having superior conductive properties can be realized. In addition, the manufacturing process is simple, and the productivity is superior.

[0094]

In addition, since this solid electrolyte 15 is formed by the polyaddition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte can be easily formed, and in addition, since being in-situ formed in a cell device, the solid electrolyte is preferably applied to a photovoltaic device or the like which uses an iodine redox couple.

[0095]

In addition, the dye-sensitized solar cell 11 described above can be formed as described below.

[0096]

First, on one major surface of the transparent substrate 2, the transparent electrode 3 is formed, and on the transparent electrode 3, the semiconductor layer 4 carrying the dye is formed, thereby forming the semiconductor electrode. In addition, on one major surface of the transparent substrate 8, the transparent electrode 7 is formed, and on the transparent electrode 7, the platinum layer 6 processed by platinum-chloride treatment is formed. Subsequently, after the semiconductor layer 4 carrying the dye and the platinum layer 6 are disposed to face each other, side surfaces of the transparent substrate 2 and the transparent substrate 7 are sealed so that the semiconductor electrode and the counter electrode are not brought into contact with each other. In this step, the distance between the semiconductor electrode and the counter electrode is not particularly limited; however, the distance is generally 1 to 100  $\mu\text{m}$  and is more preferably 1 to 30  $\mu\text{m}$ . When this distance is excessively large, the conductivity is decreased, and as a result, the amount of photocurrent is unfavorably decreased.

[0097]

In addition, although the sealing method is not

specifically limited, a material having light resistance, insulating properties, and humidity resistance is preferably used, and for example, an epoxy resin, a UV curable resin, an acrylic adhesive, ethylene vinyl acetate (EVA), a ceramic, or a heat sealable film may be used. In addition, the photovoltaic device of the present invention must have an inlet opening through which a solution before gelation is injected. The place of the inlet opening is not particularly limited as long as it is not on the semiconductor layer 4 carrying the dye and part of the counter electrode that faces the semiconductor layer 4, and hence the inlet opening may be provided at an optional place.

[0098]

Next, a mixed solution is prepared which serves as the precursor for the gel electrolyte 15. First, the first compound having at least two isocyanate groups is dissolved in an ionic liquid containing a redox couple. Subsequently, the second compound having at least two nucleophilic groups containing active hydrogen is dissolved, thereby preparing the mixed solution before gelation. In this process, the order of the dissolution of the compounds is not particularly limited; however, when the first compound and the second compound are directly mixed with each other, a reaction rapidly may occur in some cases so as to immediately produce a solid form. Hence, the case described

above is not preferable.

[0099]

Next, the mixed solution is injected into the dye-sensitized solar cell 11 assembled as described above. The method for injecting the mixed solution is not specifically limited; however, a method is preferably performed in which after the compounds are thoroughly dissolved, the mixed solution is injected inside the dye-sensitized solar cell 11 which is sealed beforehand and which is provided with the inlet opening for the mixed solution. In this case, a method in which after several droplets of the mixed solution are dripped in the inlet opening, the mixed solution is injected using a capillary phenomenon can be easily performed. In addition, whenever necessary, the injection of the mixed solution may be performed at a reduced pressure. After the mixed solution is entirely injected, a mixed solution before gelation remaining at the inlet opening is removed, and this inlet opening is then sealed. This sealing method is also not particularly limited, and whenever necessary, sealing may be performed by bonding a glass plate with a sealing agent.

[0100]

After injected into the dye-sensitized solar cell 11, the mixed solution before gelation is preferably allowed to stand still until the polymerization between the first

compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen is completed by the polyaddition reaction. The standing time is not particularly limited; however, the time in which the mixed solution injected inside the dye-sensitized solar cell 11 completely loses its viscosity and the gelation is completed is in the range of approximately 1 minute to 48 hours. However, the time described above may vary in some cases in accordance with various conditions, that is, in accordance with the selections of the first and the second compounds, the ionic liquid, and the like. In addition, the ambient temperature for the standing is not particularly limited; however, it is generally in the range of 0 to 100°C and is preferably in the range of 0 to 60°C for reducing the influence on the dye and the electrolyte.

[0101]

As described above, the dye-sensitized solar cell 11 can be formed.

[0102]

In the dye-sensitized solar cell 11 thus formed, since the solid electrolyte 15 described above is provided, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a dye-sensitized solar cell having superior

reliability can be realized.

[0103]

In addition, in this dye-sensitized solar cell 11, since the matrix polymer of the solid electrolyte 15 is a polymer formed of the above first and second compounds by the polyaddition reaction and has a chemically cross-linking structure, the solid electrolyte 15 is not changed into a liquid form by heating, and hence a dye-sensitized solar cell having superior mechanical properties and durability can be realized.

[0104]

In addition, in the case in which a solid electrolyte is used for a dye-sensitized solar cell, when the solid electrolyte is formed beforehand and is then provided so as to be in close contact with a semiconductor layer for forming the dye-sensitized solar cell, since the semiconductor layer and the solid electrolyte each having its own shape, that is, each having a solidified surface shape, are brought into contact with each other, adhesion between the semiconductor layer and the solid electrolyte is not satisfactory. As a result, due to the insufficient contact between the semiconductor layer and the solid electrolyte, a problem may arise in that the photoelectric conversion efficiency is decreased.

[0105]

However, in this dye-sensitized solar cell 11, since the solid electrolyte 15 is formed by polymerization after the precursor therefor, which is in a state before polymerization and which has fluidity, is injected into the dye-sensitized solar cell, the pores in the electrode surface are sufficiently filled with the electrolyte, and in addition, the adhesion of the solid electrolyte 15 with the semiconductor fine particles of the semiconductor layer 4 and the platinum layer 6 used as the counter electrode can be increased. Accordingly, since the contact between the semiconductor layer 4 and the solid electrolyte 15 can be satisfactorily secured, the chemical bonding condition at the electrochemical interface between the solid electrolyte 15 and the electrode surface is improved, thereby realizing a dye-sensitized solar cell having superior photoelectric conversion properties.

[0106]

In addition, in this dye-sensitized solar cell 11, since the solid electrolyte 15 is formed by polymerization using the polyaddition reaction as described above, the use of heat and/or active rays is not necessary. Accordingly, the electrolyte composition is not degraded by the use of heat and/or active rays in production of the solid electrolyte, and hence a photovoltaic device having superior photoelectric conversion properties can be realized. In

addition, the manufacturing process is simple, and the productivity is superior.

[0107]

In addition, in this dye-sensitized solar cell 11, since this solid electrolyte 15 is formed by the polyaddition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte 15 can be easily formed, and in addition, the solid electrolyte 15 can be in-situ formed in a cell device. Hence, according to this dye-sensitized solar cell 11, a dye-sensitized solar cell having superior photoelectric conversion properties can be easily and reliably formed.

[0108]

Next, another solid electrolyte and photovoltaic device, according to the present invention, and manufacturing methods thereof will be described in detail. Fig. 3 is a cross-sectional view showing the structure of another dye-sensitized solar cell formed in accordance with the present invention. As shown in Fig. 3, a dye-sensitized solar cell 21 is formed of the transparent substrate 2, the transparent electrode 3, the semiconductor layer 4, a solid electrolyte 25, the platinum layer 6 processed by platinum-chloride

treatment, the transparent electrode 7, and the transparent substrate 8. In this dye-sensitized solar cell 21, the same reference numerals of the constituent elements of the above dye-sensitized solar cell 1 shown in Fig. 1 are used to designate the same constituent elements, and descriptions thereof in detail will be omitted. A difference between the dye-sensitized solar cell 21 and the dye-sensitized solar cell 1, that is, in more particular, the solid electrolyte 25, will only be described.

[0109]

The solid electrolyte 25 is a formed of a gel electrolyte or an entire solid electrolyte, which functions as a carrier transport layer, and is a polymer material having a cross-linking matrix formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen. Accordingly, in this dye-sensitized solar cell, when the solid electrolyte 25 formed of a gel electrolyte or an entire solid electrolyte is used for the electrolyte layer as described above, the decrease of the electrolyte layer due to leakage and evaporation of liquid, which occurs when an electrolyte solution is used, can be prevented, and hence a dye-sensitized solar cell having superior cell properties and reliability can be realized.

[0110]

In this case, the unsaturated double bond of the first compound is preferably an  $\alpha$ ,  $\beta$ -unsaturated carbonyl group, an  $\alpha$ ,  $\beta$ -unsaturated sulfonyl group, or an  $\alpha$ ,  $\beta$ -unsaturated nitrile group, and among those mentioned above, the  $\alpha$ ,  $\beta$ -unsaturated carbonyl group is more preferable. The first compound and the second compound each preferably have ether, ester, carbonate, alkyl, perfluorocarbon, nitrile, tertiary amine, or the like at the main and the side chains. The number of the unsaturated double bonds of the first compound is preferably at least two; however, when the number of the nucleophilic groups containing active hydrogen of the second compound is two, the number of the unsaturated double bonds of the first compound must be at least three. As is the case described above, the number of the nucleophilic groups containing active hydrogen of the second compound is preferably at least two; however, when the number of the unsaturated double bonds of the first compound is two, the number of the nucleophilic groups containing active hydrogen of the second compound must be at least three.

[0111]

In addition, compounds having different structures may be used alone or in combination.

[0112]

When the solid electrolyte 25 is a gel electrolyte, the

gel electrolyte is composed of an electrolyte composition containing a solvent and the above cross-linking matrix, and the ratio of the cross-linking matrix is 3 to 50 percent by weight of the gel electrolyte. As the solvent forming the electrolyte composition, for example, there may be mentioned water, alcohol, ether, ester, carbonic acid ester, lactone, carboxylic acid ester, phosphate triester, heterocyclic compound, nitrile, ketone, amide, nitormethane, halogenated hydrocarbon, dimethyl sulfoxide, sulfolane, N-methyl pyrrolidone, 1,3-dimethyl imidazolidinone, 3-methyl oxazolidinone, or hydrocarbon. However, the solvent is not limited to those mentioned above, and the aforementioned solvents may be used alone or in combination. In addition, among the solvents mentioned above, aprotic nonaqueous solvents are more preferable.

[0113]

When the ratio of the electrolyte composition to the gel electrolyte is increased, the ion conductivity is increased, and on the other hand, the mechanical strength is decreased. On the contrary, as the ratio of the electrolyte composition to the gel electrolyte is decreased, the mechanical strength is increased, but the ion conductivity is decreased. Hence, the ratio of the electrolyte composition is preferably 50 to 97 percent by weight of the gel electrolyte and more preferably 80 to 95 percent by

weight thereof.

[0114]

As the electrolytes used for the gel electrolyte, in addition to the combination of  $I_2$  and a metal iodide or an organic iodide and the combination of  $Br_2$  and a metal bromide or an organic bromide, for example, metal complexes such as ferrocyanate/ferricyanate and ferrocene/ferricinium ion; sulfur compounds, such as sodium polysulfide and alkylthiol/alkyldisulfide; viologne dyes; and hydroquinone/quinone may be used.

[0115]

As cations of the above metal compounds, Li, Na, K, Mg, Ca, Cs, and the like are preferable, and as cations of the above organic compounds, quaternary ammonium compounds, such as tetraalkylammonium, pyridinium, and imidazolium, are preferable. However, the cations are not limited to those mentioned above, and the above cations may be used alone or in combination.

[0116]

Among those mentioned above, the combination of  $I_2$  and LiI or a quaternary ammonium compound such as imidazolium iodide is preferably used as the electrolyte. The concentration of the electrolyte salt is preferably 0.05 to 5 M to the solvent and is more preferably 0.2 to 1 M. The concentration of  $I_2$  or  $Br_2$  is preferably 0.0005 to 1 M and

is more preferably 0.001 to 0.1 M. In addition, in order to improve the open-circuit voltage and short-circuit current, various additives, such as 4-tert-butylpyridine and a carboxylic acid, may also be added.

[0117]

In addition, when the solid electrolyte 25 is an entire solid electrolyte, the above entire solid electrolyte is formed of the cross-linking matrix containing iodine redox. As the electrolytes, in addition to the combination of  $I_2$  and a metal iodide or an organic iodide and the combination of  $Br_2$  and a metal bromide or an organic bromide, for example, metal complexes such as ferrocyanate/ferricyanate and ferrocene/ferricinium ion; sulfur compounds, such as sodium polysulfide and alkylthiol/alkyldisulfide; viologne dyes; and hydroquinone/quinone may be used.

[0118]

As cations of the above metal compounds, Li, Na, K, Mg, Ca, Cs, and the like are preferable, and as cations of the above organic compounds, quaternary ammonium compounds, such as tetraalkylammonium, pyridinium, and imidazolium, are preferable. However, the cations are not limited to those mentioned above, and the above cations may be used alone or in combination. Among those mentioned above, the combination of  $I_2$  and LiI or a quaternary ammonium compound such as imidazolium iodide is preferably used as the

electrolyte. The concentration of the electrolyte salt is preferably 0.05 to 5 M with respect to the monomer and is more preferably 0.2 to 1 M. In addition, the concentration of I<sub>2</sub> or Br<sub>2</sub> is preferably 0.0005 to 1 M and is more preferably 0.001 to 0.1 M.

[0119]

Since the solid electrolyte 25 having the structure as described above is formed of a gel electrolyte or an entire solid electrolyte, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a highly reliable electrolyte can be realized.

[0120]

In addition, since being a polymer formed by the Michael addition reaction between the above first and the second compounds and having a cross-linking structure, the matrix polymer of the solid electrolyte 25 is not changed into a liquid form by heating, and hence superior mechanical properties and durability can be obtained.

[0121]

In addition, since this solid electrolyte 25 is formed by polymerization after a precursor therefor, which is in a state before polymerization and which has fluidity, is brought into contact with a surface on which the solid electrolyte is to be formed, that is, with an electrode

surface, the pores in the electrode surface are sufficiently filled with the electrolyte, and in addition, the chemical bonding condition at the electrochemical interface between the solid electrolyte 25 and the electrode surface is improved, thereby obtaining superior conductive properties.

[0122]

In addition, since the solid electrolyte 25 is formed by polymerization using the Michael addition reaction between the above first and the second compounds, the use of heat and/or active rays is not necessary in polymerization. Accordingly, the electrolyte composition is not degraded by heat and/or active rays in production, and hence an electrolyte having superior conductive properties can be realized. In addition, the manufacturing process is simple, and the productivity is superior.

[0123]

In addition, since this solid electrolyte 25 is formed by the Michael addition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte 25 can be easily formed, and in addition, since being in-situ formed in a cell device, the solid electrolyte is preferably applied to a photovoltaic device or the like which uses an iodine redox

couple.

[0124]

In addition, the above dye-sensitized solar cell 21 can be formed as described below.

[0125]

First, on one major surface of the transparent substrate 2, the transparent electrode 3 is formed, and on the transparent electrode 3, the semiconductor layer 4 carrying the dye is formed, thereby forming the semiconductor electrode. In addition, on one major surface of the transparent substrate 8, the transparent electrode 7 is formed, and on the transparent electrode 7, the platinum layer 6 processed by platinum-chloride treatment is formed. Subsequently, after the semiconductor layer 4 carrying the dye and the platinum layer 6 are disposed to face each other, side surfaces of the transparent substrate 2 and the transparent substrate 7 are sealed so that the semiconductor electrode and the counter electrode are not brought into contact with each other. In this step, the distance between the semiconductor electrode and the counter electrode is not particularly limited; however, the distance is generally 1 to 100  $\mu\text{m}$  and is more preferably 1 to 30  $\mu\text{m}$ . When this distance is excessively large, the conductivity is decreased, and as a result, the amount of photocurrent is unfavorably decreased.

[0126]

In addition, although the sealing method is not specifically limited, a material having light resistance, insulating properties, and humidity resistance is preferably used, and for example, an epoxy resin, a UV curable resin, an acrylic adhesive, ethylene vinyl acetate (EVA), a ceramic, or a heat sealable film may be used. In addition, the photovoltaic device of the present invention must have an inlet opening through which a solution before gelation is injected. The place of the inlet opening is not particularly limited as long as it is not on the semiconductor layer 4 carrying the dye and part of the counter electrode that faces the semiconductor layer 4, and hence the inlet opening may be provided at an optional place.

[0127]

Next, a mixed solution is prepared which serves as the precursor for the gel electrolyte 25. The first compound having at least two unsaturated double bonds is first dissolved in an electrolyte solution which is the electrolyte composition. Subsequently, the second compound having at least two nucleophilic groups containing active hydrogen is dissolved. Since the nucleophilic groups containing active hydrogen of this second compound are inactivated by reaction with carbonic acid ester or lactone, when ester, carbonic acid ester, or lactone is used as the

solvent forming the electrolyte solution, the first compound must first be dissolved, and the second compound must then be dissolved.

[0128]

Next, the mixed solution is injected into the dye-sensitized solar cell 21 assembled as described above. The method for injecting the mixed solution is not specifically limited; however, a method is preferably performed in which after the compounds are thoroughly dissolved, the mixed solution is injected inside the dye-sensitized solar cell 21 which is sealed beforehand and which is provided with the inlet opening for the mixed solution. In this case, a method in which after several droplets of the mixed solution are dripped in the inlet opening, the mixed solution is injected using a capillary phenomenon can be easily performed. In addition, whenever necessary, the injection of the mixed solution may be performed at a reduced pressure. After the mixed solution is entirely injected, a mixed solution before gellation remaining at the inlet opening is removed, and this inlet opening is then sealed. This sealing method is also not particularly limited, and whenever necessary, sealing may be performed by bonding a glass plate with a sealing agent.

[0129]

After injected into the dye-sensitized solar cell 21,

the mixed solution before gelation is preferably allowed to stand still until the polymerization between the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen is completed by the Michael addition reaction. The standing time is not particularly limited; however, the time in which the mixed solution injected inside the dye-sensitized solar cell 21 completely loses its viscosity and the gelation is completed is in the range of approximately 1 minute to 48 hours. However, the time described above may vary in some cases in accordance with various conditions, that is, in accordance with the selections of the first and the second compounds, the electrolyte, the solvent, and the like. In addition, the ambient temperature for the standing is not particularly limited; however, it is generally in the range of 0 to 100°C and is preferably in the range of 0 to 60°C for reducing the influence on the dye and the electrolyte.

[0130]

In the dye-sensitized solar cell 21 thus formed, since the solid electrolyte 25 described above is provided, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a dye-sensitized solar cell having superior reliability can be realized.

[0131]

In addition, in this dye-sensitized solar cell 21, since the matrix polymer of the solid electrolyte 25 is a polymer formed by polymerization using the Michael addition reaction between the above first and the second compounds and has a chemically cross-linking structure, the solid electrolyte 25 is not changed into a liquid form by heating, and hence a dye-sensitized solar cell having superior mechanical properties and durability can be realized.

[0132]

In addition, in the case in which a solid electrolyte is used for a dye-sensitized solar cell, when the solid electrolyte is formed beforehand and is then provided so as to be in close contact with a semiconductor layer for forming the dye-sensitized solar cell, since the semiconductor layer and the solid electrolyte each having its own shape, that is, each having a solidified surface shape, are brought into contact with each other, adhesion between the semiconductor layer and the solid electrolyte is not satisfactory. As a result, due to the insufficient contact between the semiconductor layer and the solid electrolyte, a problem may arise in that the photoelectric conversion efficiency is decreased.

[0133]

However, in this dye-sensitized solar cell 21, since

the solid electrolyte 25 is formed by polymerization after the precursor therefor, which is in a state before polymerization and which has fluidity, is injected into the dye-sensitized solar cell, the pores in the electrode surface are sufficiently filled with the electrolyte, and in addition, the adhesion of the solid electrolyte 25 with the semiconductor fine particles of the semiconductor layer 4 and the platinum layer 6 used as the counter electrode can be increased. Accordingly, since the contact between the semiconductor layer 4 and the solid electrolyte 25 can be satisfactorily secured, the chemical bonding condition at the electrochemical interface between the solid electrolyte 25 and the electrode surface is improved, thereby realizing a dye-sensitized solar cell having superior photoelectric conversion properties.

[0134]

In addition, in this dye-sensitized solar cell 21, since the solid electrolyte 25 is formed by polymerization using the Michael addition reaction, the use of heat and/or active rays is not necessary. Accordingly, the electrolyte composition is not degraded by the use of heat and/or active rays in production of the solid electrolyte, and hence a photovoltaic device having superior photoelectric conversion properties can be realized. In addition, the manufacturing process is simple, and the productivity is superior.

[0135]

In addition, in this dye-sensitized solar cell 21, since this solid electrolyte 25 is formed by the Michael addition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte 25 can be easily formed, and in addition, the solid electrolyte 25 can be in-situ formed in a cell device. Hence, according to this dye-sensitized solar cell 21, a dye-sensitized solar cell having superior photoelectric conversion properties can be easily and reliably formed.

[0136]

In the above solid electrolyte formed using the Michael addition reaction, by using an ionic liquid, that is, a molten salt, instead of the electrolyte composition containing a solvent, a gel electrolyte having no vapor pressure can also be realized.

[0137]

Next, another solid electrolyte and photovoltaic device, according to the present invention, and manufacturing methods thereof will be described in detail. Fig. 4 is a cross-sectional view showing the structure of another dye-sensitized solar cell formed in accordance with the present

invention. As shown in Fig. 4, a dye-sensitized solar cell 31 is formed of the transparent substrate 2, the transparent electrode 3, the semiconductor layer 4, a solid electrolyte 35, the platinum layer 6 processed by platinum-chloride treatment, the transparent electrode 7, and the transparent substrate 8. In this dye-sensitized solar cell 31, the same reference numerals of the constituent elements of the above dye-sensitized solar cell 1 shown in Fig. 1 are used to designate the same constituent elements, and descriptions thereof in detail will be omitted. A difference between the dye-sensitized solar cell 31 and the dye-sensitized solar cell 1, that is, in more particular, the solid electrolyte 35, will only be described.

[0138]

The solid electrolyte 35 is a formed of a gel electrolyte which functions as a carrier transport layer, and is a polymer material having a cross-linking matrix formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen. In addition, the solid electrolyte 35 is formed of an ionic liquid containing a redox couple and 3 to 50 percent by weight of the cross-linking matrix described above. Accordingly, in this dye-sensitized solar cell, when the solid electrolyte

35 formed of the gel electrolyte is used for an electrolyte layer, the decrease of the electrolyte layer due to leakage and evaporation of liquid can be prevented, which occurs when an electrolyte solution is used, and hence a dye-sensitized solar cell having superior cell properties and reliability can be realized.

[0139]

In this case, the unsaturated double bond of the first compound is preferably an  $\alpha$ ,  $\beta$ -unsaturated carbonyl group, an  $\alpha$ ,  $\beta$ -unsaturated sulfonyl group, or an  $\alpha$ ,  $\beta$ -unsaturated nitrile group, and among those mentioned above, the  $\alpha$ ,  $\beta$ -unsaturated carbonyl group is more preferable. The first compound and the second compound each preferably have ether, ester, carbonate, alkyl, perfluorocarbon, nitrile, tertiary amine, or the like at the main and the side chains. The number of the unsaturated double bonds of the first compound is preferably at least two; however, when the number of the nucleophilic groups containing active hydrogen of the second compound is two, the number of the unsaturated double bonds of the first compound must be at least three. As is the case described above, the number of the nucleophilic groups containing active hydrogen of the second compound is preferably at least two; however, when the number of the unsaturated double bonds of the first compound is two, the number of the nucleophilic groups containing active hydrogen

of the second compound must be at least three.

[0140]

In addition, compounds having different structures may be used alone or in combination.

[0141]

As the ionic liquid, for example, there may be mentioned a pyridinium salt, an imidazolium salt, or a triazorium salt; however, the ionic liquid is not limited thereto. In addition, those mentioned above may be used alone or in combination. The melting point of the ionic liquid is preferably 100°C or less, more preferably 80°C or less, and particularly preferably 60°C or less.

[0142]

When the ratio of the ionic liquid to the gel electrolyte is increased, the ion conductivity is increased thereby, but the mechanical strength is decreased. On the other hand, as the ratio of the ionic liquid to the gel electrolyte is decreased, the mechanical strength is increased; however, the ion conductivity is decreased. Hence, the ratio of the ionic liquid is preferably 50 to 97 percent by weight of the gel electrolyte and is more preferably 80 to 95 percent by weight thereof.

[0143]

As the redox couple used for the gel electrolyte of the present invention, the combination of I<sub>2</sub> and an ionic liquid

of iodide or the combination of Br<sub>2</sub> and an ionic liquid of bromide is preferable, and in particular, the combination of I<sub>2</sub> and the ionic liquid is preferable. The concentration of the redox species to the total electrolyte is preferably 0.1 to 20 percent by weight and is more preferably 0.2 to 5 percent by weight.

[0144]

In addition, besides the materials mentioned above, as an counter anion of the ionic liquid, for example, there may be mentioned a halogen ion, such as Cl<sup>-</sup> or Br<sup>-</sup>, NSC<sup>-</sup>, BF<sub>4</sub><sup>-</sup>, PF<sub>6</sub><sup>-</sup>, ClO<sub>4</sub><sup>-</sup>, (CF<sub>3</sub>SO<sub>2</sub>)<sub>2</sub>N<sup>-</sup>, (CF<sub>3</sub>CF<sub>2</sub>SO<sub>2</sub>)<sub>2</sub>N<sup>-</sup>, CF<sub>3</sub>SO<sub>3</sub><sup>-</sup>, CF<sub>3</sub>COO<sup>-</sup>, Ph<sub>4</sub>B<sup>-</sup>, (CF<sub>3</sub>SO<sub>2</sub>)<sub>3</sub>C<sup>-</sup>, or F(HF)<sub>n</sub>, and among those mentioned above, (CF<sub>3</sub>SO<sub>2</sub>)<sub>2</sub>N<sup>-</sup> or BF<sub>4</sub><sup>-</sup> is preferable.

[0145]

In addition, as the redox couples in this case, besides the combination with a metal iodide or an organic iodide, and the combination of Br<sub>2</sub> and a metal bromide or an organic bromide, for example, metal complexes such as ferrocyanate/ferricyanate and ferrocene/ferricinium ion; sulfur compounds, such as sodium polysulfide and alkylthiol-alkyldisulfide; viologne dyes; and hydroquinone/quinone may be used.

[0146]

As cations of the above metal compounds, Li, Na, K, Mg, Ca, Cs, and the like are preferable, and as cations of the

above organic compounds, quaternary ammonium compounds, such as tetraalkylammonium, pyridinium, and imidazolium, are preferable. However, the cations are not limited to those mentioned above, and the above cations may be used alone or in combination. Among those mentioned above, an electrolyte formed in combination of  $I_2$  and a quaternary ammonium compound such as imidazolium iodide is preferably used.

[0147]

The concentration of the electrolyte salt is preferably 0.05 to 5 M with respect to the ionic liquid and is more preferably 0.2 to 1 M. The concentration of  $I_2$  or  $Br_2$  is preferably 0.0005 to 1 M and is more preferably 0.001 to 0.1 M. In addition, in order to improve the open-circuit voltage and short-circuit current, various additives, such as 4-tert-butylpyridine and a carboxylic acid, may also be added.

[0148]

Since the solid electrolyte 35 having the structure as described above is formed of the gel electrolyte, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a highly reliable electrolyte can be realized.

[0149]

In addition, since being a polymer formed by polymerization using the Michael addition reaction between

the above first and the second compounds and having a chemically cross-linking structure, the matrix polymer of this 35 solid electrolyte is not changed into a liquid form by heating, and hence superior mechanical properties and durability can be obtained.

[0150]

In addition, since the solid electrolyte 35 is formed by polymerization after a precursor therefor, which is in a state before polymerization and has fluidity, is brought into contact with a surface, such as an electrode surface, on which the solid electrolyte is to be formed, the pores in the electrode surface are sufficiently filled with the electrolyte, and in addition, the chemical bonding condition at the electrochemical interface between the solid electrolyte and the electrode surface is improved, thereby obtaining superior conductive properties.

[0151]

In addition, since the solid electrolyte 35 is formed by polymerization using the Michael addition reaction between the above first and second compounds, the use of heat and/or active rays is not necessary in polymerization. Accordingly, the electrolyte composition is not degraded by heat and/or active rays in production, and hence an electrolyte having superior conductive properties can be realized. In addition, the manufacturing process is simple,

and the productivity is superior.

[0152]

In addition, since this solid electrolyte 35 is formed by the Michael addition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte can be easily formed, and in addition, since being in-situ formed in a cell device, the solid electrolyte can be preferably applied to a photovoltaic device or the like which uses an iodine redox couple.

[0153]

In addition, the above dye-sensitized solar cell 31 can be formed as described below.

[0154]

First, on one major surface of the transparent substrate 2, the transparent electrode 3 is formed, and on the transparent electrode 3, the semiconductor layer 4 carrying the dye is formed, thereby forming the semiconductor electrode. In addition, on one major surface of the transparent substrate 8, the transparent electrode 7 is formed, and on the transparent electrode 7, the platinum layer 6 processed by platinum-chloride treatment is formed. Subsequently, the semiconductor layer 4 carrying the dye and

the platinum layer 6 are disposed to face each other, and side surfaces of the transparent substrate 2 and the transparent substrate 7 are sealed so that the semiconductor electrode and the counter electrode are not brought into contact with each other. In this step, the distance between the semiconductor electrode and the counter electrode is not particularly limited; however, the distance is generally 1 to 100  $\mu\text{m}$  and is more preferably 1 to 30  $\mu\text{m}$ . When this distance is excessively large, the conductivity is decreased, and as a result, the amount of photocurrent is unfavorably decreased.

[0155]

In addition, although the sealing method is not specifically limited, a material having light resistance, insulating properties, and humidity resistance is preferably used, and for example, an epoxy resin, a UV curable resin, an acrylic adhesive, ethylene vinyl acetate (EVA), a ceramic, or a heat sealable film may be used. In addition, the photovoltaic device of the present invention must have an inlet opening through which a solution before gelation is injected. The place of the inlet opening is not particularly limited as long as it is not on the semiconductor layer 4 carrying the dye and part of the counter electrode that faces the semiconductor layer 4, and hence the inlet opening may be provided at an optional place.

[0156]

Next, a mixed solution is prepared which serves as a precursor for the gel electrolyte 35. First, the first compound having at least two unsaturated double bonds is dissolved in an ionic liquid containing a redox couple. Subsequently, the second compound having at least two nucleophilic groups containing active hydrogen is dissolved, thereby preparing the mixed solution before gelation. In this process, the order of the dissolution of the compounds is not particularly limited; however, when the first compound and the second compound are directly mixed with each other, a reaction rapidly may occur in some cases so as to immediately produce a solid form. Hence, the case described above is not preferable.

[0157]

Next, the mixed solution is injected into the dye-sensitized solar cell 31 assembled as described above. The method for injecting the mixed solution is not specifically limited; however, a method is preferably performed in which after the compounds are thoroughly dissolved, the mixed solution is injected inside the dye-sensitized solar cell 31 which is sealed beforehand and which is provided with the inlet opening for the mixed solution. In this case, a method in which after several droplets of the mixed solution are dripped in the inlet opening, the mixed solution is

injected using a capillary phenomenon can be easily performed. In addition, whenever necessary, the injection of the mixed solution may be performed at a reduced pressure. After the mixed solution is entirely injected, a mixed solution before gelation remaining at the inlet opening is removed, and this inlet opening is then sealed. This sealing method is also not particularly limited, and whenever necessary, sealing may be performed by bonding a glass plate with a sealing agent.

[0158]

After injected into the dye-sensitized solar cell 31, the mixed solution before gelation is preferably allowed to stand still until the polymerization between the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen is completed by the Michael addition reaction. The standing time is not particularly limited; however, the time in which the mixed solution injected inside the dye-sensitized solar cell 31 completely loses its viscosity and the gelation is completed is in the range of approximately 1 minute to 48 hours. However, the time described above may vary in some cases in accordance with various conditions, that is, in accordance with the selections of the first and the second compounds, the ionic liquid, and the like. In addition, the ambient temperature

for the standing is not particularly limited; however, it is generally in the range of 0 to 100°C and is preferably in the range of 0 to 60°C for reducing the influence on the dye and the electrolyte.

[0159]

As described above, the dye-sensitized solar cell 31 can be formed.

[0160]

In the dye-sensitized solar cell 31 thus formed, since the solid electrolyte 35 described above is provided, for example, degradation in properties due to leakage and evaporation of an electrolyte solution does not occur, and hence a dye-sensitized solar cell having superior reliability can be realized.

[0161]

In addition, in this dye-sensitized solar cell 31, since the matrix polymer of the solid electrolyte 35 is a polymer formed by polymerization using the Michael addition reaction between the above first and the second compounds and has a chemically cross-linking structure, the solid electrolyte 35 is not changed into a liquid form by heating, and hence a dye-sensitized solar cell having superior mechanical properties and durability can be realized.

[0162]

In addition, in the case in which a solid electrolyte

is used for a dye-sensitized solar cell, when the solid electrolyte is formed beforehand and is then provided so as to be in close contact with a semiconductor layer for forming the dye-sensitized solar cell, since the semiconductor layer and the solid electrolyte each having its own shape, that is, each having a solidified surface shape, are brought into contact with each other, adhesion between the semiconductor layer and the solid electrolyte is not satisfactory. As a result, due to the insufficient contact between the semiconductor layer and the solid electrolyte, a problem may arise in that the photoelectric conversion efficiency is decreased.

[0163]

However, in this dye-sensitized solar cell 31, since the solid electrolyte 35 is formed by polymerization after the precursor therefor, which is in a state before polymerization and which has fluidity, is injected into the dye-sensitized solar cell, the pores in the electrode surface are sufficiently filled with the electrolyte, and in addition, the adhesion of the solid electrolyte 35 with the semiconductor fine particles of the semiconductor layer 4 and the platinum layer 6 used as the counter electrode can be increased. Accordingly, since the contact between the semiconductor layer 4 and the solid electrolyte 35 can be satisfactorily secured, the chemical bonding condition at

the electrochemical interface between the solid electrolyte 35 and the electrode surface is improved, thereby realizing a dye-sensitized solar cell having superior photoelectric conversion properties.

[0164]

In addition, in this dye-sensitized solar cell 31, since the solid electrolyte 35 is formed by polymerization using the Michael addition reaction as described above, the use of heat and/or active rays is not necessary. Accordingly, the electrolyte composition is not degraded by the use of heat and/or active rays in production of the solid electrolyte, and hence a photovoltaic device having superior photoelectric conversion properties can be realized. In addition, the manufacturing process is simple, and the productivity is superior.

[0165]

In addition, in this dye-sensitized solar cell 31, since this solid electrolyte 35 is formed by the Michael addition reaction and is not formed by a radical polymerization method, even when iodine, which functions as a polymerization inhibitor in a radical polymerization method, is contained in the electrolyte composition, the solid electrolyte 35 can be easily formed, and in addition, the solid electrolyte 35 can be in-situ formed in a cell device. Hence, according to this dye-sensitized solar cell

31, a dye-sensitized solar cell having superior photoelectric conversion properties can be easily and reliably formed.

[0166]

In addition, the photovoltaic device of the present invention is not limited to the structures described above and may be optionally formed in various shapes in accordance with applications.

[0167]

[Examples]

Hereinafter, the present invention will be described in more detail with reference to particular experimental results.

[0168]

<Experiment 1>

In Experiment 1, a photovoltaic device provided with a gel electrolyte of the present invention was investigated.

[0169]

[Example 1]

In Example 1, as the photovoltaic device of the present invention, a dye-sensitized solar cell was formed as described below. First, a  $\text{TiO}_2$  paste was formed. The formation of the  $\text{TiO}_2$  paste was performed with reference to the information disclosed in "Novel Technique of Dye-Sensitized Solar Cell" (CMC). First, 125 ml of titanium

isopropoxide was slowly dripped into 750 ml of an aqueous nitric acid solution at a concentration of 0.1 M while it was stirred at room temperature. After the dripping was finished, the mixture thus formed was placed in a temperature-control oven at 80°C and was then stirred for 8 hours, thereby forming a white translucent sol solution.

[0170]

Next, after this sol solution was allowed to stand still for cooling until the temperature thereof decreased to room temperature and was then filtrated using a glass filter, "mess-up" was performed to obtain 700 ml of a solution. The sol solution thus obtained was placed in an autoclave and was then processed by hydrothermal treatment at 220°C for 12 hours, followed by performing dispersion process by ultrasonic treatment for 1 hour.

[0171]

Next, this solution was concentrated using an evaporator at 40°C so that the content of  $\text{TiO}_2$  was adjusted to 11 percent by weight. Next, polyethylene oxide (PEO) having a molecular weight of 500,000 was added to this concentrated sol solution and was then uniformly mixed together by a planetary ball mill, and as a result, a  $\text{TiO}_2$  paste having an increased viscosity was obtained.

[0172]

Next, after the  $\text{TiO}_2$  paste thus obtained was applied by

a screen printing method onto a fluorine-doped conductive glass substrate having a sheet resistance of  $10 \Omega/\square$  as a transparent electrode so as to form a film having a size of 1 cm by 1 cm,  $\text{TiO}_2$  was sintered on the conductive glass under the conditions in which a temperature of  $450^\circ\text{C}$  was maintained for 30 minutes, and as a result, a semiconductor layer was formed.

[0173]

Next, the semiconductor layer thus formed was immersed for 12 hours in a dehydrated ethanol solution containing 0.5 mM of cis-bis(isothiocyanate)-N,N-bis(2,2'-dipyridyl-4,4'-dicarboxylic acid)-ruthenium (II) dihydrate and 20 mM of deoxycholic acid, so that a dye was adsorbed on the semiconductor layer. Next, washing was sequentially performed using an ethanol solution containing 4-tert-butylpyridine and dehydrated ethanol in that order, and drying was performed at a dark place, thereby forming a semiconductor electrode.

[0174]

Next, platinum was sputtered to have a thickness of 100 nm onto a fluorine-doped conductive glass substrate which was provided with an inlet opening having a diameter of 1 mm and which had a sheet resistance of  $10 \Omega/\square$ , and several droplets of an ethanol solution containing chloroplatinic acid were dripped onto the platinum thus sputtered, followed

by heating to 450°C, thereby forming a counter electrode.

[0175]

Next, the semiconductor layer of the semiconductor electrode and the platinum surface of the counter electrode, which were formed as described above, were disposed to face each other, and sealing was performed using a heat sealable film having a thickness of 20  $\mu\text{m}$  so that the above two electrodes were not brought into contact with each other, thereby forming a dye-sensitized solar cell.

[0176]

In addition, in 3 g of a solvent composed of ethylene carbonate and propylene carbonate at a weight ratio of 1 to 1, 0.2 g of lithium iodide (LiI), 0.479 g of 1-propyl-2,3-dimethylimidazolium iodide, 0.0381 g of iodine ( $\text{I}_2$ ), and 0.2 g of 4-tert-butylpyridine were dissolved, thereby preparing an electrolyte solution which was an electrolyte composition.

[0177]

Next, after 100 ppm of dibutyl tin dilaurate was added as a catalyst to the above electrolyte solution, polyoxyethylene glycerin (molecular weight of 1,200) was mixed therewith as the second compound having at least two nucleophilic groups containing active hydrogen, and subsequently, tolylene diisocyanate was also mixed as the first compound having at least two isocyanate groups, so that a mixed solution was prepared. The mixing ratio on a

molar basis of the isocyanate group to the nucleophilic group containing active hydrogen was 1 to 1, and the polyoxyethylene glycerin and the tolylene diisocyanate were mixed so that the total thereof was 6 percent by weight with respect to the electrolyte solution.

[0178]

Next, several droplets of the above mixed solution were dripped in the inlet opening of the dye-sensitized solar cell prepared beforehand, and the mixed solution was then injected therein by evacuation, followed by sealing of the inlet opening using a heat sealable film and a glass plate. After the mixed solution was injected, this dye-sensitized solar cell was allowed to stand still at 80°C for 30 minutes, and a dye-sensitized solar cell was obtained which is provided with a gel electrolyte formed by gelation of the electrolyte composition.

[0179]

[Example 2]

A dye-sensitized solar cell was formed in the same manner as that in Example 1 except that polyoxyethylene pentaerythritol (molecular weight of 1,600) was used as the second compound.

[0180]

[Example 3]

A dye-sensitized solar cell was formed in the same

manner as that in Example 1 except that hexamethylene diisocyanate was used as the first compound.

[0181]

[Example 4]

A dye-sensitized solar cell was formed in the same manner as that in Example 1 except that isophorone diisocyanate was used as the first compound.

[0182]

[Example 5]

A dye-sensitized solar cell was formed in the same manner as that in Example 1 except that glycerin was used as the second compound and that the catalyst was not added.

[0183]

[Comparative Example 1]

A dye-sensitized solar cell was formed in the same manner as that in Example 1 except that the first compound, the second compound, and the catalyst were not mixed with the electrolyte solution which was the electrolyte composition and that the electrolyte solution was used as an electrolyte layer.

[0184]

[Comparative Example 2]

A dye-sensitized solar cell was formed in the same manner as that in Example 1 except that a mixed solution was prepared by mixing polyethylene glycol triacrylate with the

electrolyte solution so as to be 6 percent by weight thereto and by mixing 0.15 percent by weight of tert-butyl peroxydipivalate as a radical polymerization initiator, and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0185]

[Comparative Example 3]

A dye-sensitized solar cell was formed in the same manner as that in Example 1 except that a mixed solution was prepared by mixing polyethylene glycol diacrylate with the electrolyte solution so as to be 6 percent by weight thereto and by mixing 0.15 percent by weight of tert-butyl peroxydipivalate as a radical polymerization initiator, and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0186]

[Comparative Example 4]

After a mixed solution was prepared in the same manner as that in Example 1, several droplets of the mixed solution were dripped onto the TiO<sub>2</sub> film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 1 except that the PET film

was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0187]

[Comparative Example 5]

After a mixed solution was prepared in the same manner as that in Example 2, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 2 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0188]

[Comparative Example 6]

After a mixed solution was prepared in the same manner as that in Example 3, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the

same manner as that in Example 3 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0189]

[Comparative Example 7]

After a mixed solution was prepared in the same manner as that in Example 4, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 4 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0190]

[Comparative Example 8]

After a mixed solution was prepared in the same manner as that in Example 5, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape.

Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 5 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0191]

The preparation conditions of the mixed solutions according to the examples and comparative examples described above are shown in Table 1.

[0192]

[Table 1]

	First compound	Second compound	Catalyst
Example 1	tolylene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 2	tolylene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Example 3	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 4	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 5	tolylene 2,4-diisocyanate	glycerin	No addition
Comparative example 1	No addition	No addition	No addition
Comparative example 2	polyethylene glycol triacrylate	--	tert-butyl peroxyphthalate
Comparative example 3	polyethylene glycol diacrylate	--	tert-butyl peroxyphthalate
Comparative example 4	tolylene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 5	tolylene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Comparative example 6	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 7	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 8	tolylene 2,4-diisocyanate	glycerin	No addition

[0193]

The photoelectric conversion efficiencies of the dye-sensitized solar cells of Examples 1 to 5 and Comparative Examples 1 to 8 formed as described above were evaluated. The photoelectric conversion efficiency was measured as described below.

[0194]

Measurement of photoelectric conversion efficiency

The measurement of the photoelectric conversion efficiency was performed by the steps of connecting alligator clips to the fluorine-doped conductive glass substrate of the dye-sensitized solar cell at the semiconductor electrode side and the fluorine-doped conductive glass substrate at the counter electrode side, irradiating the dye-sensitized solar cell with artificial-sunlight (AM1.5,  $100 \text{ mW/cm}^2$ ), and measuring a current generated thereby. The results are shown in Table 2.

[0195]

[Table 2]

	Gelation	Photoelectric conversion efficiency (%)
Example 1	○	7.9
Example 2	○	7.7
Example 3	○	7.9
Example 4	○	7.8
Example 5	○	7.5
Comparative example 1	-	8.2
Comparative example 2	×	4.5 <sup>*1</sup>
Comparative example 3	×	4.7 <sup>*1</sup>
Comparative example 4	○	5.6
Comparative example 5	○	5.6
Comparative example 6	○	5.5
Comparative example 7	○	5.4
Comparative example 8	○	5.3

\*1 indicates a reference value since no gelation occurred.

[0196]

In addition, the individual dye-sensitized solar cells were each decomposed after the measurement of the photoelectric conversion efficiency so that an electrolyte portion provided inside the solar cell was recovered, and the occurrence of gelation of the electrolyte was confirmed by visual inspection. The results are also shown in Table 2. In Table 2, ○ indicates that the gelation occurs, and × indicates that the gelation does not occur.

[0197]

As can be seen from Table 2, in Comparative Examples 2 and 3 in which a conventional radical polymerization method was used, it is understood that the electrolyte composition is not in-situ gelled in the dye-sensitized solar cell. On the other hand, in Examples 1 to 5 in which the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized by the polyaddition reaction, it is understood that the electrolyte composition is in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0198]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized by the polyaddition reaction, the electrolyte composition can be in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0199]

In addition, when Examples 1 to 5 are compared with Comparative Example 1, the decrease in photoelectric conversion efficiency of the dye-sensitized solar cell having the gel electrolyte of the present invention is small

as compared to that of the dye-sensitized solar cell formed of the electrolyte solution, and a superior photoelectric conversion efficiency is shown; hence, it can be said that a superior dye-sensitized solar cell is realized.

[0200]

In addition, as can be seen from Table 2, in Comparative Examples 4 to 8 in which the gel electrolyte was formed beforehand on the semiconductor layer by polymerization using the polyaddition reaction between the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen and in which the dye-sensitized solar cell was formed by providing the counter electrode on the gel electrolyte thus formed, it is understood that although the gel electrolyte is obtained, the photoelectric conversion efficiency is considerably decreased as compared to that of the dye-sensitized solar cell formed using the electrolyte solution. The reason for this is believed that since the chemical bonding condition at the electrochemical interface between the gel electrolyte which is formed beforehand by gelation and the electrode surface is not satisfactory, the resistance is increased.

[0201]

On the other hand, in Examples 1 to 5 according to the present invention in which the first compound having at

least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized in the dye-sensitized solar cell by the polyaddition reaction so that gelation was in-situ performed, it is understood that the decrease in photoelectric conversion efficiency is small as compared to that of the dye-sensitized solar cell formed using the electrolyte solution and that hence a superior photoelectric conversion efficiency is obtained. The reason for this is believed that since the mixed solution is gelled while being in contact with the electrode surface, the chemical bonding condition at the electrochemical interface between the gel electrolyte and the electrode surface is improved, and hence the resistance is decreased.

[0202]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized in the dye-sensitized solar cell by the polyaddition reaction, the electrolyte composition can be in-situ gelled in the dye-sensitized solar cell so as to form the gel electrolyte. Accordingly, it is said that when the gel electrolyte described above is used, a dye-sensitized solar cell having a superior photoelectric conversion efficiency can be

realized.

[0203]

<Experiment 2>

In Experiment 2, a photovoltaic device provided with an entire solid electrolyte of the present invention was investigated.

[0204]

[Example 6]

In Example 6, a dye-sensitized solar cell was formed in the same manner as that in Example 1 except that a mixed solution was prepared as described below.

[0205]

Preparation of Mixed Solution

After 0.2 g of lithium iodide (LiI), 0.479 g of 1-propyl-2,3-dimethylimidazolium iodide, 0.0381 g of iodine (I<sub>2</sub>), and 0.2 g of 4-tert-butylpyridine were dissolved in 3 g of polyoxyethylene glycerin (molecular weight of 1,200) used as the second compound having at least two nucleophilic groups containing active hydrogen, 100 ppm of dibutyl tin dilaurate was added as a catalyst.

[0206]

Next, tolylene diisocyanate was mixed as the first compound having at least two isocyanate groups. The mixing ratio on a molar basis of the isocyanate group to the nucleophilic group containing active hydrogen was set to 1

to 1. Accordingly, the mixed solution was prepared.

[0207]

[Example 7]

A dye-sensitized solar cell was formed in the same manner as that in Example 6 except that polyoxyethylene pentaerythritol (molecular weight of 1,600) was used as the second compound.

[0208]

[Example 8]

A dye-sensitized solar cell was formed in the same manner as that in Example 6 except that hexamethylene diisocyanate was used as the first compound.

[0209]

[Example 9]

A dye-sensitized solar cell was formed in the same manner as that in Example 6 except that isophorone diisocyanate was used as the first compound.

[0210]

[Example 10]

A dye-sensitized solar cell was formed in the same manner as that in Example 6 except that glycerin was used as the second compound and that the catalyst was not added.

[0211]

[Comparative Example 9]

A dye-sensitized solar cell was formed in the same

manner as that in Example 6 except that a mixed solution was prepared by using polyethylene glycol triacrylate and by mixing 0.15 percent by weight of tert-butyl peroxy-pivalate as a radical polymerization initiator and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0212]

[Comparative Example 10]

A dye-sensitized solar cell was formed in the same manner as that in Example 6 except that a mixed solution was prepared by using polyethylene glycol diacrylate and by mixing 0.15 percent by weight of tert-butyl peroxy-pivalate as a radical polymerization initiator and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0213]

[Comparative Example 11]

After a mixed solution was prepared in the same manner as that in Example 6, several droplets of the mixed solution were dripped onto the TiO<sub>2</sub> film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 6 except that the PET film

was removed after solidification, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0214]

[Comparative Example 12]

After a mixed solution was prepared in the same manner as that in Example 7, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 7 except that the PET film was removed after solidification, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0215]

[Comparative Example 13]

After a mixed solution was prepared in the same manner as that in Example 8, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the

same manner as that in Example 8 except that the PET film was removed after solidification, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0216]

[Comparative Example 14]

After a mixed solution was prepared in the same manner as that in Example 9, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 9 except that the PET film was removed after solidification, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0217]

[Comparative Example 15]

After a mixed solution was prepared in the same manner as that in Example 10, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid

electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 10 except that the PET film was removed after solidification, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0218]

The preparation conditions of the mixed solutions according to the examples and comparative examples described above are shown in Table 3.

[0219]

[Table 3]

	First compound	Second compound	Catalyst
Example 6	tolyene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 7	tolyene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Example 8	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 9	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 10	tolyene 2,4-diisocyanate	glycerin	No addition
Comparative example 9	polyethylene glycol triacrylate	--	tert-butyl peroxyphthalate
Comparative example 10	polyethylene glycol diacrylate	--	tert-butyl peroxyphthalate
Comparative example 11	tolyene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 12	tolyene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Comparative example 13	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 14	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 15	tolyene 2,4-diisocyanate	glycerin	No addition

[0220]

The photoelectric conversion efficiencies of the dye-sensitized solar cells of Examples 6 to 10 and Comparative Examples 9 to 15 formed as described above were evaluated in the same manner as that described above. The results are shown in Table 4.

[0221]

[Table 4]

	Gelation	Photoelectric conversion efficiency (%)
Example 6	○	1.5
Example 7	○	1.4
Example 8	○	1.5
Example 9	○	1.3
Example 10	○	1.0
Comparative example 9	×	0.5* <sup>2</sup>
Comparative example 10	×	0.5* <sup>2</sup>
Comparative example 11	○	0.7
Comparative example 12	○	0.7
Comparative example 13	○	0.7
Comparative example 14	○	0.6
Comparative example 15	○	0.5

\*2 indicates a reference value since no gelation occurred.

[0222]

In addition, the individual dye-sensitized solar cells

were each decomposed after the measurement of the photoelectric conversion efficiency so that an electrolyte portion provided inside the solar cell was recovered, and the occurrence of entire solidification of the electrolyte was confirmed by visual inspection. The results are also shown in Table 4. In Table 4, O indicates that the entire solidification occurs, and x indicates that the entire solidification does not occur.

[0223]

As can be seen from Table 4, in Comparative Examples 9 and 10 in which a conventional radical polymerization method was used, it is understood that the electrolyte composition is not entirely in-situ solidified in the dye-sensitized solar cell. On the other hand, in Examples 6 to 10 in which the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized by the polyaddition reaction, it is understood that the electrolyte composition is entirely in-situ solidified in the dye-sensitized solar cell and that the entire solid electrolyte is obtained.

[0224]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups

containing active hydrogen are polymerized by the polyaddition reaction, the electrolyte composition can be entirely in-situ solidified in the dye-sensitized solar cell and that the entire solid electrolyte can be obtained.

[0225]

In addition, as can be seen from Table 4, in Examples 6 to 10 according to the present invention in which the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized in the dye-sensitized solar cell by the polyaddition reaction so that the entire solidification was in-situ performed, it is understood that a high photoelectric conversion efficiency can be obtained as compared to that obtained in Comparative Examples 11 to 15 in which the dye-sensitized solar cell was formed by the steps of forming the entire solid electrolyte beforehand on the semiconductor layer, and providing the counter electrode on the entire solid electrolyte.

[0226]

The reason a high photoelectric conversion efficiency is obtained in Examples 6 to 10 is believed that, since the mixed solution is entirely solidified while being in contact with the electrode surface, the chemical bonding condition at the electrochemical interface between the entire solid electrolyte and the electrode surface is improved, and hence

the resistance is decreased. On the other hand, the reason the photoelectric conversion efficiency is low in Comparative Examples 11 to 15 is believed that, since the chemical bonding condition at the electrochemical interface between the electrode surface and the entire solid electrolyte which is entirely solidified beforehand is not satisfactory, the resistance is increased.

[0227]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized in the dye-sensitized solar cell by the polyaddition reaction, the electrolyte composition can be entirely in-situ solidified in the dye-sensitized solar cell so as to form the entire solid electrolyte. Accordingly, it is said that when the entire solid electrolyte described above is used, a dye-sensitized solar cell having a superior photoelectric conversion efficiency can be realized.

[0228]

<Experiment 3>

In Experiment 3, a photovoltaic device provided with another gel electrolyte of the present invention was investigated.

[0229]

[Example 21]

In Example 21, a dye-sensitized solar cell was formed in the same manner as that in Example 1 except that a mixed solution was prepared as described below to form a gel electrolyte.

[0230]

First, in 1-ethyl-3-methylimidazolium (hereinafter referred to as "EMI" in some cases)-bis(trifluoromethanesulfonyl)imide (TFSI), 0.9 M of 1-propyl-2,3-dimethylimidazolium iodide, 30 mM of iodine ( $I_2$ ), and 0.5 M of 4-tert-butylpyridine were dissolved, so that an ionic liquid containing a redox couple was prepared.

[0231]

Next, after 100 ppm of a catalyst, dibutyl tin dilaurate, was added to the ionic liquid, polyoxyethylene glycerin (molecular weight of 1,200) was mixed therewith as the second compound having at least two nucleophilic groups containing active hydrogen, and tolylene diisocyanate was further mixed as the first compound having at least two isocyanate groups, thereby preparing the mixed solution. The mixing ratio on a molar basis of the isocyanate group to the nucleophilic group containing active hydrogen was set to 1 to 1, and the polyoxyethylene glycerin and the tolylene diisocyanate were mixed so that the total thereof is 6 percent by weight with respect to the ionic liquid.

[0232]

After several droplets of the mixed solution were dripped in the inlet opening of the device assembled beforehand, by evacuation in an atmosphere at a temperature of 50°C, the mixed solution was injected into the device, and subsequently, the inlet opening was sealed with a heat sealable film and a glass plate. After the injection, this device was allowed to stand still for 12 hours at room temperature, and hence a dye-sensitized solar cell provided with a gel electrolyte was obtained, the gel electrolyte being formed by gelation of the ionic liquid containing a redox couple.

[0233]

[Example 22]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that polyoxyethylene pentaerythritol (molecular weight of 1,600) was used as the second compound.

[0234]

[Example 23]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that hexamethylene diisocyanate was used as the first compound.

[0235]

[Example 24]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that isophorone diisocyanate was used as the first compound.

[0236]

[Example 25]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that glycerin was used as the second compound and that the catalyst was not added.

[0237]

[Example 26]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that EMI-iodide was used for the ionic liquid, and as the redox species, 30 mM of iodine ( $I_2$ ) and 0.5 M of 4-tert-butylpyridine were dissolved so that an ionic liquid containing a redox couple was prepared.

[0238]

[Example 27]

A dye-sensitized solar cell was formed in the same manner as that in Example 26 except that polyoxyethylene pentaerythritol (molecular weight of 1,600) was used as the second compound.

[0239]

[Example 28]

A dye-sensitized solar cell was formed in the same

manner as that in Example 26 except that hexamethylene diisocyanate was used as the first compound.

[0240]

[Example 29]

A dye-sensitized solar cell was formed in the same manner as that in Example 26 except that isophorone diisocyanate was used as the first compound.

[0241]

[Example 30]

A dye-sensitized solar cell was formed in the same manner as that in Example 26 except that glycerin was used as the second compound and that the catalyst was not added.

[0242]

[Comparative Example 21]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that the first compound, the second compound, and the catalyst were not mixed with the ionic liquid.

[0243]

[Comparative Example 22]

A dye-sensitized solar cell was formed in the same manner as that in Example 26 except that the first compound, the second compound, and the catalyst were not mixed with the ionic liquid.

[0244]

[Comparative Example 23]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol triacrylate and 0.15 percent by weight of tert-butyl peroxyvalate as a radical polymerization initiator with respect to the ionic liquid of Example 21 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0245]

[Comparative Example 24]

A dye-sensitized solar cell was formed in the same manner as that in Example 21 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol diacrylate and 0.15 percent by weight of tert-butyl peroxyvalate as a radical polymerization initiator with respect to the ionic liquid of Example 21 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0246]

[Comparative Example 25]

A dye-sensitized solar cell was formed in the same manner as that in Example 26 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol triacrylate and 0.15 percent by weight of tert-butyl

peroxypivalate as a radical polymerization initiator with respect to the ionic liquid of Example 26 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0247]

[Comparative Example 26]

A dye-sensitized solar cell was formed in the same manner as that in Example 26 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol diacrylate and 0.15 percent by weight of tert-butyl peroxypivalate as a radical polymerization initiator with respect to the ionic liquid of Example 26 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0248]

[Comparative Example 27]

After a mixed solution was prepared in the same manner as that in Example 21, several droplets of the mixed solution were dripped onto the TiO<sub>2</sub> film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 21 except that the PET film was removed after gelation, the counter

electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0249]

[Comparative Example 28]

After a mixed solution was prepared in the same manner as that in Example 22, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 22 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0250]

[Comparative Example 29]

After a mixed solution was prepared in the same manner as that in Example 23, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 23 except

that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0251]

[Comparative Example 30]

After a mixed solution was prepared in the same manner as that in Example 24, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 24 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0252]

[Comparative Example 31]

After a mixed solution was prepared in the same manner as that in Example 25, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell

was formed in the same manner as that in Example 25 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0253]

[Comparative Example 32]

After a mixed solution was prepared in the same manner as that in Example 26, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 26 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0254]

[Comparative Example 33]

After a mixed solution was prepared in the same manner as that in Example 27, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer

in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 27 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0255]

[Comparative Example 34]

After a mixed solution was prepared in the same manner as that in Example 28, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 28 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0256]

[Comparative Example 35]

After a mixed solution was prepared in the same manner as that in Example 29, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand

still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 29 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0257]

[Comparative Example 36]

After a mixed solution was prepared in the same manner as that in Example 30, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 30 except that the PET film was removed after gelation, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0258]

The preparation conditions of the mixed solutions according to the examples and comparative examples described above are shown in Table 5.

[0259]

[Table 5]

	First compound	Second compound	Catalyst
Example 21	tolylene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 22	tolylene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Example 23	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 24	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 25	tolylene 2,4-diisocyanate	glycerin	No addition
Example 26	tolylene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 27	tolylene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Example 28	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 29	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Example 30	tolylene 2,4-diisocyanate	glycerin	No addition
Comparative example 21	No addition	No addition	No addition
Comparative example 22	No addition	No addition	No addition
Comparative example 23	polyethylene glycol triacrylate	—	tert-butyl peroxy-pivalate
Comparative example 24	polyethylene glycol diacrylate	—	tert-butyl peroxy-pivalate
Comparative example 25	polyethylene glycol triacrylate	—	tert-butyl peroxy-pivalate
Comparative example 26	polyethylene glycol diacrylate	—	tert-butyl peroxy-pivalate
Comparative example 27	tolylene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate

Comparative example 28	tolylene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Comparative example 29	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 30	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 31	tolylene 2,4-diisocyanate	glycerin	No addition
Comparative example 32	tolylene 2,4-diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 33	tolylene 2,4-diisocyanate	polyoxyethylene pentaerythritol (molecular weight of 1,600)	dibutyl tin dilaurate
Comparative example 34	hexamethylene diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 35	isophorone diisocyanate	polyoxyethylene glycerin (molecular weight of 1,200)	dibutyl tin dilaurate
Comparative example 36	tolylene 2,4-diisocyanate	glycerin	No addition

[0260]

The photoelectric conversion efficiencies of the dye-sensitized solar cells of Examples 21 to 30 and Comparative Examples 21 to 36 formed as described above were evaluated in the same manner as that described above. The results are shown in Table 6.

[0261]

[Table 6]

	Gelation	Photoelectric conversion efficiency (%)
Example 21	○	1.3
Example 22	○	1.4
Example 23	○	1.3
Example 24	○	1.2
Example 25	○	1.0
Example 26	○	1.3
Example 27	○	1.3
Example 28	○	1.2
Example 29	○	1.2
Example 30	○	0.9
Comparative example 21	×	1.5 <sup>*3</sup>
Comparative example 22	×	1.3 <sup>*3</sup>
Comparative example 23	×	0.7 <sup>*3</sup>
Comparative example 24	×	0.6 <sup>*3</sup>
Comparative example 25	×	0.4 <sup>*3</sup>
Comparative example 26	×	0.4 <sup>*3</sup>
Comparative example 27	○	0.8
Comparative example 28	○	0.8
Comparative example 29	○	0.7
Comparative example 30	○	0.7
Comparative example 31	○	0.5
Comparative example 32	○	0.7
Comparative example 33	○	0.7
Comparative example 34	○	0.6
Comparative example 35	○	0.6
Comparative example 36	○	0.5

\*3 indicates a reference value since no gelation occurred.

[0262]

In addition, the individual dye-sensitized solar cells were each decomposed after the measurement of the photoelectric conversion efficiency so that an electrolyte portion provided inside the solar cell was recovered, and the occurrence of gelation of the electrolyte was confirmed

by visual inspection. The results are also shown in Table 6. In Table 6, O indicates that the gelation occurs, and x indicates that the gelation does not occur.

[0263]

As can be seen from Table 6, in Comparative Examples 23 to 26 in which a conventional radical polymerization method was used, it is understood that the ionic liquid containing a redox couple is not in-situ gelled in the dye-sensitized solar cell. On the other hand, in Examples 21 to 30 in which the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized by the polyaddition reaction, it is understood that the ionic liquid containing a redox couple is in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0264]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized by the polyaddition reaction, the ionic liquid containing a redox couple can be in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0265]

In addition, when Examples 21 to 30 are compared with Comparative Examples 21 and 22, the decrease in photoelectric conversion efficiency of the dye-sensitized solar cell having the gel electrolyte of the present invention is small as compared to that of the dye-sensitized solar cell formed of the ionic liquid containing a redox couple, and a superior photoelectric conversion efficiency is shown; hence, it can be said that a superior dye-sensitized solar cell is realized.

[0266]

In addition, as can be seen from Table 6, in Comparative Examples 27 to 36 in which the gel electrolyte was formed beforehand on the semiconductor layer by polymerization using the polyaddition reaction between the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen and in which the dye-sensitized solar cell was formed by providing the counter electrode on the gel electrolyte thus formed, it is understood that although the gel electrolyte is obtained, compared to the dye-sensitized solar cell formed using the ionic liquid containing a redox couple, the photoelectric conversion efficiency is considerably decreased. The reason for this is believed that since the chemical bonding condition at the electrochemical interface between the gel electrolyte which

is formed beforehand by gelation and the electrode surface is not satisfactory, the resistance is increased.

[0267]

On the other hand, in Examples 21 to 30 according to the present invention in which the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized in the dye-sensitized solar cell by the polyaddition reaction so that gelation was in-situ performed, it is understood that the decrease in photoelectric conversion efficiency is small as compared to that of the dye-sensitized solar cell formed of the ionic liquid containing a redox couple and that hence a superior photoelectric conversion efficiency is obtained. The reason for this is believed that since the mixed solution is gelled while being in contact with the electrode surface, the chemical bonding condition at the electrochemical interface between the gel electrolyte and the electrode surface is improved, and hence the resistance is decreased.

[0268]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized in the dye-sensitized solar cell by the polyaddition reaction, the

ionic liquid containing a redox couple can be in-situ gelled in the dye-sensitized solar cell so as to form the gel electrolyte. Accordingly, it is said that when the gel electrolyte described above is used, a dye-sensitized solar cell having a superior photoelectric conversion efficiency can be realized.

[0269]

<Experiment 4>

In Experiment 4, a photovoltaic device provided with another gel electrolyte of the present invention was investigated.

[0270]

[Example 41]

In Example 41, a dye-sensitized solar cell was formed in the same manner as that in Example 1 except that a mixed solution was prepared as described below.

[0271]

In 3 g of a solvent composed of ethylene carbonate and propylene carbonate at a mixing ratio of 1 to 1 on a weight basis, 0.2 g of lithium iodide (LiI), 0.479 g of 1-propyl-2,3-dimethylimidazolium iodide, 0.0381 g of iodine (I<sub>2</sub>), and 0.2 g of 4-tert-butylpyridine were dissolved, so that an electrolyte solution which was an electrolyte composition was prepared.

[0272]

Next, polyethylene glycol triacrylate (molecular weight of 3,000) as the first compound having at least two unsaturated double bonds was mixed with the electrolyte solution, and 4,4'-trimethylenedipiperidine was further mixed as the second compound having at least two nucleophilic groups containing active hydrogen, thereby preparing the mixed solution. The mixing ratio on a molar basis of the unsaturated double bond to the nucleophilic group containing active hydrogen was set to 1 to 1, and the polyethylene glycol triacrylate and the 4,4'-trimethylenedipiperidine were mixed so that the total thereof was 6 percent by weight to the electrolyte solution.

[0273]

After several droplets of the mixed solution were dripped in the inlet opening of the device assembled beforehand, evacuation was performed so that the mixed solution was injected into the device, and subsequently, the inlet opening was sealed with a heat sealable film and a glass plate. After the injection, this device was allowed to stand still for 12 hours at room temperature, and hence a dye-sensitized solar cell provided with a gel electrolyte was obtained, the gel electrolyte being formed by gelation of the electrolyte composition.

[0274]

[Example 42]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that polyethylene glycol diacrylate was used as the first compound and that polyethylene glycol tripiperidine was used as the second compound.

[0275]

[Example 43]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that polyethylene glycol tripiperidine was used as the second compound.

[0276]

[Example 44]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that polyethylene glycol triamine was used as the second compound.

[0277]

[Example 45]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that polyethyleneimine (molecular weight of 600) was used as the second compound.

[0278]

[Comparative Example 41]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that the first compound and the second compound were not mixed with the electrolyte

solution which was the electrolyte composition and that a liquid electrolyte composition was used as an electrolyte layer.

[0279]

[Comparative Example 42]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that a mixed solution was prepared by mixing polyethylene glycol triacrylate as the first compound so as to be 6 percent by weight with respect to the electrolyte solution and by mixing 0.15 percent by weight of tert-butyl peroxy-pivalate as a radical polymerization initiator without mixing the second compound and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0280]

[Comparative Example 43]

A dye-sensitized solar cell was formed in the same manner as that in Example 41 except that the mixed solution was prepared by mixing polyethylene glycol diacrylate as the first compound so as to be 6 percent by weight with respect to the electrolyte solution and by mixing 0.15 percent by weight of tert-butyl peroxy-pivalate as a radical polymerization initiator without mixing the second compound and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0281]

[Comparative Example 44]

After a mixed solution was prepared in the same manner as that in Example 41, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 41 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0282]

[Comparative Example 45]

After a mixed solution was prepared in the same manner as that in Example 42, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 42 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte

layer, and sealing was performed using a heat sealable film.

[0283]

[Comparative Example 46]

After a mixed solution was prepared in the same manner as that in Example 43, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 43 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0284]

[Comparative Example 47]

After a mixed solution was prepared in the same manner as that in Example 44, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 44 except that after the gelation, the PET film was removed, the

counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0285]

[Comparative Example 48]

After a mixed solution was prepared in the same manner as that in Example 45, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 45 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0286]

The preparation conditions of the mixed solutions according to the examples and comparative examples described above are shown in Table 7.

[0287]

[Table 7]

	First compound	Second compound
Example 41	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Example 42	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Example 43	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Example 44	polyethylene glycol triacrylate	polyethylene glycol triamine
Example 45	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)
Comparative example 41	No addition	No addition
Comparative example 42	polyethylene glycol triacrylate	--
Comparative example 43	polyethylene glycol diacrylate	--
Comparative example 44	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Comparative example 45	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Comparative example 46	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Comparative example 47	polyethylene glycol triacrylate	polyethylene glycol triamine
Comparative example 48	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)

[0288]

The photoelectric conversion efficiencies of the dye-sensitized solar cells of Examples 41 to 45 and Comparative

Examples 41 to 48 formed as described above were evaluated in the same manner as that described above. The results are shown in Table 8.

[0289]

[Table 8]

	Gelation	Photoelectric conversion efficiency (%)
Example 41	○	7.8
Example 42	○	7.9
Example 43	○	8.0
Example 44	○	8.0
Example 45	○	7.8
Comparative example 41	-	8.2
Comparative example 42	×	4.5 <sup>*4</sup>
Comparative example 43	×	4.7 <sup>*4</sup>
Comparative example 44	○	5.8
Comparative example 45	○	5.7
Comparative example 46	○	5.9
Comparative example 47	○	5.9
Comparative example 48	○	5.6

\*4 indicates a reference value since no gelation occurred.

[0290]

In addition, the individual dye-sensitized solar cells were each decomposed after the measurement of the photoelectric conversion efficiency so that an electrolyte portion provided inside the solar cell was recovered, and

the occurrence of gelation of the electrolyte was confirmed by visual inspection. The results are also shown in Table 8. In Table 8, O indicates that gelation occurs, and x indicates that gelation does not occur.

[0291]

As can be seen from Table 8, in Comparative Examples 42 and 43 in which a conventional radical polymerization method was used, it is understood that the electrolyte composition is not in-situ gelled in the dye-sensitized solar cell. On the other hand, in Examples 41 to 45 in which the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized by the Michael addition reaction, it is understood that the electrolyte composition is in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0292]

From the above results, it is understood that when the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized by the Michael addition reaction, the electrolyte composition can be in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0293]

In addition, when Examples 41 to 45 are compared with Comparative Example 41, the decrease in photoelectric conversion efficiency of the dye-sensitized solar cell having the gel electrolyte of the present invention is small as compared to that of the dye-sensitized solar cell formed of the liquid electrolyte composition, that is, the electrolyte solution, and a superior photoelectric conversion efficiency is shown; hence, it can be said that a superior dye-sensitized solar cell is realized.

[0294]

In addition, as can be seen from Table 8, in Comparative Examples 44 to 48 in which the gel electrolyte was formed beforehand on the semiconductor layer by polymerization using the Michael addition reaction between the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen and in which the dye-sensitized solar cell was formed by providing the counter electrode on the gel electrolyte thus formed, it is understood that although the gel electrolyte is obtained, compared to the dye-sensitized solar cell formed using the liquid electrolyte composition, that is, the electrolyte solution, the photoelectric conversion efficiency is considerably decreased. The reason for this is believed that since the chemical bonding condition at the

electrochemical interface between the gel electrolyte which is formed beforehand by gelation and the electrode surface is not satisfactory, the resistance is increased.

[0295]

On the other hand, in Examples 41 to 45 according to the present invention in which the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized in the dye-sensitized solar cell by the Michael addition reaction so that gelation was in-situ performed, it is understood that the decrease in photoelectric conversion efficiency is small as compared to that of the dye-sensitized solar cell formed using the liquid electrolyte composition, that is, the electrolyte solution, and that hence a superior photoelectric conversion efficiency is obtained. The reason for this is believed that since the mixed solution is gelled while being in contact with the electrode surface, the chemical bonding condition at the electrochemical interface between the gel electrolyte and the electrode surface is improved, and hence the resistance is decreased.

[0296]

From the above results, it is understood that when the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic

groups containing active hydrogen are polymerized in the dye-sensitized solar cell by the Michael addition reaction, the electrolyte composition can be in-situ gelled in the dye-sensitized solar cell so as to form the gel electrolyte. Accordingly, it is said that when the gel electrolyte described above is used, a dye-sensitized solar cell having a superior photoelectric conversion efficiency can be realized.

[0297]

<Experiment 5>

In Experiment 5, a photovoltaic device provided with another entire solid electrolyte of the present invention was investigated.

[0298]

[Example 46]

In Example 46, a dye-sensitized solar cell was formed in the same manner as that in Example 41 except that a mixed solution was prepared as described below.

[0299]

In 3 g of polyethylene glycol triacrylate (molecular weight of 3,000) used as the first compound having at least two unsaturated double bonds, 0.2 g of lithium iodide (LiI), 0.479 g of 1-propyl-2,3-dimethylimidazolium iodide, 0.0381 g of iodine (I<sub>2</sub>), and 0.2 g of 4-tert-butylpyridine were dissolved.

[0300]

Next, 4,4'-trimethylenedipiperidine was mixed which was used as the second compound having at least two nucleophilic groups containing active hydrogen. The mixing ratio on a molar basis of the unsaturated double bond to the nucleophilic group containing active hydrogen was set to 1 to 1. Accordingly, the mixed solution was prepared.

[0301]

[Example 47]

A dye-sensitized solar cell was formed in the same manner as that in Example 46 except that polyethylene glycol diacrylate was used as the first compound and that polyethylene glycol tripiperidine was used as the second compound.

[0302]

[Example 48]

A dye-sensitized solar cell was formed in the same manner as that in Example 46 except that polyethylene glycol tripiperidine was used as the second compound.

[0303]

[Example 49]

A dye-sensitized solar cell was formed in the same manner as that in Example 46 except that polyethylene glycol triamine was used as the second compound.

[0304]

[Example 50]

A dye-sensitized solar cell was formed in the same manner as that in Example 46 except that polyethyleneimine (molecular weight of 600) was used as the second compound.

[0305]

[Comparative Example 49]

A dye-sensitized solar cell was formed in the same manner as that in Example 46 except that a mixed solution was prepared by mixing polyethylene glycol triacrylate as the first compound so as to be 6 percent by weight with respect to the electrolyte composition and by mixing 0.15 percent by weight of tert-butyl peroxyvalate as a radical polymerization initiator without mixing the second compound and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0306]

[Comparative Example 50]

A dye-sensitized solar cell was formed in the same manner as that in Example 46 except that a mixed solution was prepared by mixing polyethylene glycol diacrylate as the first compound so as to be 6 percent by weight with respect to the electrolyte composition and by mixing 0.15 percent by weight of tert-butyl peroxyvalate as a radical polymerization initiator without mixing the second compound and that after injected, this mixed solution was allowed to

stand still at 60°C for 1 hour.

[0307]

[Comparative Example 51]

After a mixed solution was prepared in the same manner as that in Example 46, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 46 except that after the solidification, the PET film was removed, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0308]

[Comparative Example 52]

After a mixed solution was prepared in the same manner as that in Example 47, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example

47 except that after the solidification, the PET film was removed, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0309]

[Comparative Example 53]

After a mixed solution was prepared in the same manner as that in Example 48, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 48 except that after the solidification, the PET film was removed, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0310]

[Comparative Example 54]

After a mixed solution was prepared in the same manner as that in Example 49, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand

still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 49 except that after the solidification, the PET film was removed, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0311]

[Comparative Example 55]

After a mixed solution was prepared in the same manner as that in Example 50, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming an entire solid electrolyte in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 50 except that after the solidification, the PET film was removed, the counter electrode was then provided on the entire solid electrolyte, and sealing was performed using a heat sealable film.

[0312]

The preparation conditions of the mixed solutions according to the examples and comparative examples described above are shown in Table 9.

[0313]

[Table 9]

	First compound	Second compound
Example 46	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Example 47	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Example 48	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Example 49	polyethylene glycol triacrylate	polyethylene glycol triamine
Example 50	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)
Comparative example 49	polyethylene glycol triacrylate	--
Comparative example 50	polyethylene glycol diacrylate	--
Comparative example 51	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Comparative example 52	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Comparative example 53	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Comparative example 54	polyethylene glycol triacrylate	polyethylene glycol triamine
Comparative example 55	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)

[0314]

The photoelectric conversion efficiencies of the dye-sensitized solar cells of Examples 46 to 50 and Comparative

Examples 49 to 55 formed as described above were evaluated in the same manner as described above. The results are shown in Table 10.

[0315]

[Table 10]

	Gelation	Photoelectric conversion efficiency (%)
Example 46	○	1.5
Example 47	○	1.6
Example 48	○	1.7
Example 49	○	1.7
Example 50	○	1.6
Comparative example 49	×	0.5 <sup>*5</sup>
Comparative example 50	×	0.5 <sup>*5</sup>
Comparative example 51	○	0.8
Comparative example 52	○	0.8
Comparative example 53	○	0.7
Comparative example 54	○	0.7
Comparative example 55	○	0.5

\*5 indicates a reference value since no gelation occurred.

[0316]

In addition, the individual dye-sensitized solar cells were each decomposed after the measurement of the photoelectric conversion efficiency so that an electrolyte portion provided inside the solar cell was recovered, and the occurrence of entire solidification of the electrolyte

composition was confirmed by visual inspection. The results are also shown in Table 10. In Table 10, O indicates that the entire solidification occurs, and x indicates that the entire solidification does not occur.

[0317]

As can be seen from Table 10, in Comparative Examples 42 and 43 in which a conventional radical polymerization method was used, it is understood that the electrolyte composition is not entirely in-situ solidified in the dye-sensitized solar cell. On the other hand, in Examples 46 to 50 in which the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized by the Michael addition reaction, it is understood that the electrolyte composition is in-situ gelled in the dye-sensitized solar cell and that the entire solid electrolyte is obtained.

[0318]

From the above results, it is understood that when the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized by the Michael addition reaction, the electrolyte composition can be entirely in-situ solidified in the dye-sensitized solar cell and that the entire solid electrolyte can be obtained.

[0319]

In addition, as can be seen from Table 10, in Examples 46 to 50 according to the present invention in which the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized in the dye-sensitized solar cell by the Michael addition reaction so that the entire solidification was in-situ performed, it is understood that a high photoelectric conversion efficiency can be obtained as compared to that obtained in Comparative Examples 51 to 55 in which the dye-sensitized solar cell was formed by the steps of forming the entire solid electrolyte beforehand on the semiconductor layer, and providing the counter electrode on the entire solid electrolyte.

[0320]

In Examples 46 to 50, the reason a high photoelectric conversion efficiency is obtained is believed that since the electrolyte solution is entirely solidified while being in contact with the electrode surface, the chemical bonding condition at the electrochemical interface between the entire solid electrolyte and the electrode surface is improved, and hence the resistance is decreased. On the other hand, in Comparative Examples 51 to 55, the reason the photoelectric conversion efficiency is low is believed that

since the chemical bonding condition at the electrochemical interface between the electrode surface and the entire solid electrolyte which is entirely solidified beforehand is not satisfactory, the resistance is increased.

[0321]

From the above results, it is understood that when the first compound having at least two isocyanate groups and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized by the Michael addition reaction in the dye-sensitized solar cell, the electrolyte composition can be entirely in-situ solidified in the dye-sensitized solar cell so as to form the entire solid electrolyte. Accordingly, it is said that when the entire solid electrolyte described above is used, a dye-sensitized solar cell having a superior photoelectric conversion efficiency can be realized.

[0322]

<Experiment 6>

In Experiment 6, a photovoltaic device provided with another gel electrolyte of the present invention was investigated.

[0323]

[Example 51]

In Example 51, a dye-sensitized solar cell was formed in the same manner as that in Example 1 except that a gel

electrolyte was formed by preparing an ionic liquid as described below.

[0324]

First, in 1-ethyl-3-methylimidazolium (hereinafter referred to as "EMI")-bis(trifluoromethanesulfonyl)imide (TFSI), 0.9 M of 1-propyl-2,3-dimethylimidazolium iodide, 30 mM of iodine ( $I_2$ ), and 0.5 M of 4-tert-butylpyridine were dissolved, so that an ionic liquid containing a redox couple was prepared.

[0325]

Next, polyethylene glycol triacrylate (molecular weight of 3,000) as the first compound having at least two unsaturated double bonds was mixed with the ionic liquid, and 4,4'-trimethylenedipiperidine was then mixed as the second compound having at least two nucleophilic groups containing active hydrogen, thereby preparing a mixed solution. The mixing ratio on a molar basis of the unsaturated double bond to the nucleophilic group containing active hydrogen was set to 1 to 1, and the polyethylene glycol triacrylate and the 4,4'-trimethylenedipiperidine were mixed so that the total thereof is 6 percent by weight with respect to the ionic liquid.

[0326]

After several droplets of the mixed solution were dripped in the inlet opening of the device assembled

beforehand, evacuation was performed in an atmosphere at a temperature of 50°C so that the mixed solution was injected into the device, and subsequently, the inlet opening was sealed with a heat sealable film and a glass plate. After the injection, this device was allowed to stand still for 12 hours at room temperature, and as a result, a dye-sensitized solar cell provided with the gel electrolyte was obtained, the gel electrolyte being formed by gelation of the ionic liquid containing a redox couple.

[0327]

[Example 52]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that polyethylene glycol diacrylate was used as the first compound and that polyethylene glycol tripiperidine was used as the second compound.

[0328]

[Example 53]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that polyethylene glycol tripiperidine was used as the second compound.

[0329]

[Example 54]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that polyethylene glycol

triamine was used as the second compound.

[0330]

[Example 55]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that polyethyleneimine (molecular weight of 600) was used as the second compound.

[0331]

[Example 56]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that EMI-iodide was used for the ionic liquid, and 30 mM of iodine ( $I_2$ ) as the redox species and 0.5 M of 4-tert-butylpyridine were dissolved therein so that an ionic liquid containing a redox couple was prepared.

[0332]

[Example 57]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that polyethylene glycol diacrylate was used as the first compound and that polyethylene glycol tri Piperidine was used as the second compound.

[0333]

[Example 58]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that polyethylene glycol

tripiperidine was used as the second compound.

[0334]

[Example 59]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that polyethylene glycol triamine was used as the second compound.

[0335]

[Example 60]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that polyethyleneimine (molecular weight of 600) was used as the second compound.

[0336]

[Comparative Example 51]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that the first compound and the second compound were not mixed with the ionic liquid.

[0337]

[Comparative Example 52]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that the first compound and the second compound were not mixed with the ionic liquid.

[0338]

[Comparative Example 53]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that a mixed solution

was prepared by mixing 6 percent by weight of polyethylene glycol triacrylate and 0.15 percent by weight of tert-butyl peroxydivalate as a radical polymerization initiator with respect to the ionic liquid of Example 51 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0339]

[Comparative Example 54]

A dye-sensitized solar cell was formed in the same manner as that in Example 51 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol diacrylate and 0.15 percent by weight of tert-butyl peroxydivalate as a radical polymerization initiator with respect to the ionic liquid of Example 51 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0340]

[Comparative Example 55]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol triacrylate and 0.15 percent by weight of tert-butyl peroxydivalate as a radical polymerization initiator with respect to the ionic liquid of Example 56 and that after injected, this mixed solution was allowed to stand still at

60°C for 1 hour.

[0341]

[Comparative Example 56]

A dye-sensitized solar cell was formed in the same manner as that in Example 56 except that a mixed solution was prepared by mixing 6 percent by weight of polyethylene glycol diacrylate and 0.15 percent by weight of tert-butyl peroxyvalate as a radical polymerization initiator with respect to the ionic liquid of Example 56 and that after injected, this mixed solution was allowed to stand still at 60°C for 1 hour.

[0342]

[Comparative Example 57]

After a mixed solution was prepared in the same manner as that in Example 51, several droplets of the mixed solution were dripped onto the TiO<sub>2</sub> film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 51 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0343]

[Comparative Example 58]

After a mixed solution was prepared in the same manner as that in Example 52, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 52 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0344]

[Comparative Example 59]

After a mixed solution was prepared in the same manner as that in Example 53, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 53 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0345]

[Comparative Example 60]

After a mixed solution was prepared in the same manner as that in Example 54, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 54 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0346]

[Comparative Example 61]

After a mixed solution was prepared in the same manner as that in Example 55, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 55 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte

layer, and sealing was performed using a heat sealable film.

[0347]

[Comparative Example 62]

After a mixed solution was prepared in the same manner as that in Example 56, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 56 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0348]

[Comparative Example 63]

After a mixed solution was prepared in the same manner as that in Example 57, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 57 except that after the gelation, the PET film was removed, the

counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0349]

[Comparative Example 64]

After a mixed solution was prepared in the same manner as that in Example 58, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 58 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0350]

[Comparative Example 65]

After a mixed solution was prepared in the same manner as that in Example 59, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 59 except

that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0351]

[Comparative Example 66]

After a mixed solution was prepared in the same manner as that in Example 60, several droplets of the mixed solution were dripped onto the  $\text{TiO}_2$  film which was the semiconductor layer, and a polyethylene terephthalate (PET) film was then provided thereon and was allowed to stand still for 12 hours, thereby forming a gel electrolyte layer in a film shape. Subsequently, a dye-sensitized solar cell was formed in the same manner as that in Example 60 except that after the gelation, the PET film was removed, the counter electrode was then provided on the gel electrolyte layer, and sealing was performed using a heat sealable film.

[0352]

The preparation conditions of the mixed solutions according to the examples and comparative examples described above are shown in Table 11.

[0353]

[Table 11]

	First compound	Second compound
Example 51	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Example 52	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Example 53	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Example 54	polyethylene glycol triacrylate	polyethylene glycol triamine
Example 55	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)
Example 56	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Example 57	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Example 58	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Example 59	polyethylene glycol triacrylate	polyethylene glycol triamine
Example 60	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)
Comparative example 51	No addition	No addition
Comparative example 52	No addition	No addition
Comparative example 53	polyethylene glycol triacrylate	--
Comparative example 54	polyethylene glycol diacrylate	--
Comparative example 55	polyethylene glycol triacrylate	--
Comparative example 56	polyethylene glycol diacrylate	--
Comparative example 57	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Comparative example 58	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Comparative example 59	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Comparative example 60	polyethylene glycol triacrylate	polyethylene glycol triamine
Comparative example 61	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)
Comparative example 62	polyethylene glycol triacrylate	4,4'-trimethylenedipiperidine
Comparative example 63	polyethylene glycol diacrylate	polyethylene glycol tripiperidine
Comparative example 64	polyethylene glycol triacrylate	polyethylene glycol tripiperidine
Comparative example 65	polyethylene glycol triacrylate	polyethylene glycol triamine
Comparative example 66	polyethylene glycol triacrylate	polyethyleneimine (molecular weight of 600)

[0354]

The photoelectric conversion efficiencies of the dye-sensitized solar cells of Examples 51 to 60 and Comparative Examples 51 to 66 formed as described above were evaluated in the same manner as described above. The results are shown in Table 12.

[0355]

[Table 12]

	Gelation	Photoelectric conversion efficiency (%)
Example 51	O	1.2
Example 52	O	1.3
Example 53	O	1.3
Example 54	O	1.2
Example 55	O	1.1
Example 56	O	1.1
Example 57	O	1.2
Example 58	O	1.2
Example 59	O	1.1
Example 60	O	1.0
Comparative example 51	x	1.5 <sup>*6</sup>
Comparative example 52	x	1.3 <sup>*6</sup>
Comparative example 53	x	0.7 <sup>*6</sup>
Comparative example 54	x	0.6 <sup>*6</sup>
Comparative example 55	x	0.4 <sup>*6</sup>
Comparative example 56	x	0.4 <sup>*6</sup>
Comparative example 57	O	0.8
Comparative example 58	O	0.8
Comparative example 59	O	0.8
Comparative example 60	O	0.7
Comparative example 61	O	0.6
Comparative example 62	O	0.8
Comparative example 63	O	0.8
Comparative example 64	O	0.7
Comparative example 65	O	0.7
Comparative example 66	O	0.5

\*6 indicates a reference value since no gelation occurred.

[0356]

In addition, the individual dye-sensitized solar cells were each decomposed after the measurement of the photoelectric conversion efficiency so that an electrolyte portion provided inside the solar cell was recovered, and the occurrence of gelation of the electrolyte was confirmed by visual inspection. The results are also shown in Table 12. In Table 12, O indicates that gelation occurs, and x indicates that gelation does not occur.

[0357]

As can be seen from Table 12, in Comparative Examples 53 to 56 in which a conventional radical polymerization method was used, it is understood that the ionic liquid containing a redox couple is not in-situ gelled in the dye-sensitized solar cell. On the other hand, in Examples 51 to 60 in which the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized by the Michael addition reaction, it is understood that the ionic liquid containing a redox couple is in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0358]

From the above results, it is understood that when the first compound having at least two unsaturated double bonds

and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized by the Michael addition reaction, the ionic liquid containing a redox couple can be in-situ gelled in the dye-sensitized solar cell and that the gel electrolyte is obtained.

[0359]

In addition, when Examples 51 to 60 are compared with Comparative Examples 51 and 52, the decrease in photoelectric conversion efficiency of the dye-sensitized solar cell having the gel electrolyte of the present invention is small as compared to that of the dye-sensitized solar cell formed of the ionic liquid containing a redox couple, and a superior photoelectric conversion efficiency is shown; hence, it can be said that superior dye-sensitized solar cell is realized.

[0360]

In addition, as can be seen from Table 12, in Comparative Examples 57 to 66 in which the gel electrolyte was formed beforehand on the semiconductor layer by polymerization using the Michael addition reaction between the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen and in which the dye-sensitized solar cell was formed by providing the counter electrode on the gel electrolyte thus formed, it is

understood that although the gel electrolyte is obtained, compared to the dye-sensitized solar cell formed using the ionic liquid containing a redox couple, the photoelectric conversion efficiency is considerably decreased. The reason for this is believed that since the chemical bonding condition at the electrochemical interface between the gel electrolyte which is formed beforehand by gelation and the electrode surface is not satisfactory, the resistance is increased.

[0361]

On the other hand, in Examples 51 to 60 according to the present invention in which the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen were polymerized in the dye-sensitized solar cell by the Michael addition reaction so that gelation was in-situ performed, it is understood that the decrease in photoelectric conversion efficiency is small as compared to that of the dye-sensitized solar cell formed using the ionic liquid containing a redox couple and that hence a superior photoelectric conversion efficiency is obtained. The reason for this is believed that since the mixed solution is gelled while being in contact with the electrode surface, the chemical bonding condition at the electrochemical interface between the gel electrolyte and the electrode surface is

improved, and hence the resistance is decreased.

[0362]

From the above results, it is understood that when the first compound having at least two unsaturated double bonds and the second compound having at least two nucleophilic groups containing active hydrogen are polymerized in the dye-sensitized solar cell by the Michael addition reaction, the ionic liquid containing a redox couple can be in-situ gelled in the dye-sensitized solar cell so as to form the gel electrolyte. Accordingly, it is said that when the gel electrolyte described above is used, a dye-sensitized solar cell having a superior photoelectric conversion efficiency can be realized.

[0363]

[Advantages]

The solid electrolyte of the present invention is a solid electrolyte having an electrolyte composition and a matrix polymer, and the matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen. In addition, the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0364]

In addition, another solid electrolyte of the present invention is a solid electrolyte having an electrolyte composition and a matrix polymer, and the matrix polymer is formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen. In addition, the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0365]

Since the solid electrolyte having the structure as described above is formed by polymerization after the precursor therefor, which is in a state before polymerization and has fluidity, is brought into contact with the surface, such as an electrode surface, on which the solid electrolyte is to be formed, the solid electrolyte can be in-situ formed, for example, in a cell device, and hence the chemical bonding condition at the electrochemical interface between the solid electrolyte and the electrode surface is improved. In addition, since heat and/or active rays are not necessary in polymerization, the electrolyte composition is not degraded by heat and/or active rays in production.

[0366]

In addition, since being a polymer formed by polymerization using a polyaddition reaction between the above first and the second compounds and having a chemically cross-linking structure, the matrix polymer of this solid electrolyte is not changed into a liquid form by application of heat, and hence superior mechanical properties and durability can be obtained.

[0367]

Hence, according to the solid electrolyte of the present invention, an electrolyte having superior conductive properties can be realized.

[0368]

In addition, a photovoltaic device using the solid electrolyte of the present invention is a photovoltaic device which has an electrolyte layer and a semiconductor layer composed of semiconductor particles carrying a dye, the layers being provided between a counter electrode and an electrode formed on a surface of a transparent substrate. The electrolyte layer has a redox couple, an electrolyte composition, and a matrix polymer, and the matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen. In addition,

the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0369]

In addition, a method for manufacturing this photovoltaic device is a method for manufacturing a photovoltaic device which has an electrolyte layer and a semiconductor layer composed of semiconductor particles carrying a dye, the layers being provided between a counter electrode and an electrode formed on a surface of a transparent substrate. In this method, after the photovoltaic device is assembled, a mixed solution including a first compound having at least two isocyanate groups, a second compound having at least two nucleophilic groups containing active hydrogen, and an electrolyte composition containing a redox couple is injected into the photovoltaic device, and the first compound and the second compound are polymerized in the photovoltaic device by a polyaddition reaction for solidification, so that the above electrolyte layer is formed.

[0370]

In addition, another photovoltaic device of the present invention is a photovoltaic device which has an electrolyte layer and a semiconductor layer composed of semiconductor particles carrying a dye, the layers being provided between

a counter electrode and an electrode formed on a surface of a transparent substrate. The electrolyte layer has a redox couple, an electrolyte composition, and a matrix polymer, and the matrix polymer is formed by polymerization using a Michael addition reaction between a first compound having at least two unsaturated double bonds and a second compound having at least two nucleophilic groups containing active hydrogen. In addition, the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[0371]

In addition, a method for manufacturing this photovoltaic device is a method for manufacturing a photovoltaic device which has a solid electrolyte layer and a semiconductor layer composed of semiconductor particles carrying a dye, the layers being provided between a counter electrode and an electrode formed on a surface of a transparent substrate. In this method, after the photovoltaic device is assembled, a mixed solution including a first compound having at least two unsaturated double bonds, a second compound having at least two nucleophilic groups containing active hydrogen, and an electrolyte composition containing a redox couple is injected into the photovoltaic device, and the first compound and the second

compound are polymerized in the photovoltaic device by a Michael addition reaction for solidification, so that the solid electrolyte layer is formed.

[0372]

According to the photovoltaic device of the present invention described above, the mixed solution which is in a state before polymerization and which has fluidity is injected into the photovoltaic device and is then polymerized, thereby forming the solid electrolyte of the present invention. Hence, the solid electrolyte can be in-situ formed in a cell device, and as a result, a photovoltaic device having superior photoelectric conversion properties can be easily and reliably realized.

[Brief Description of the Drawings]

[Fig. 1]

Fig. 1 is a cross-sectional view showing one structural example of a dye-sensitized solar cell formed according to the present invention.

[Fig. 2]

Fig. 2 is a cross-sectional view showing another structural example of a dye-sensitized solar cell formed according to the present invention.

[Fig. 3]

Fig. 3 is a cross-sectional view showing another structural example of a dye-sensitized solar cell formed

according to the present invention.

[Fig. 4]

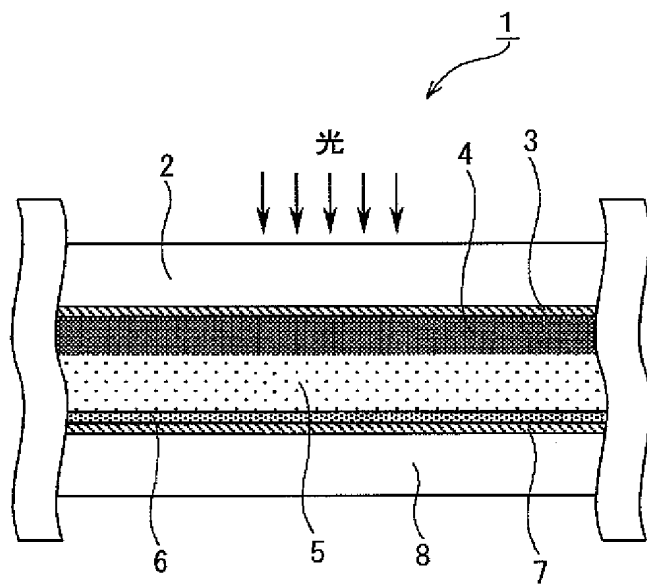
Fig. 4 is a cross-sectional view showing another structural example of a dye-sensitized solar cell formed according to the present invention.

[Reference Numerals]

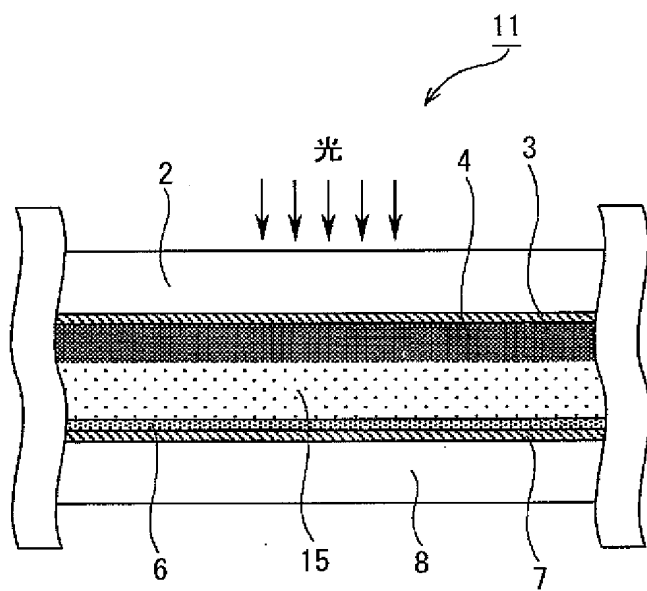
- 1, 11, 21, 31 dye-sensitive solar cell
- 2 transparent substrate
- 3 transparent electrode
- 4 semiconductor layer
- 5, 15, 25, 35 solid electrolyte
- 6 platinum electrode
- 7 transparent electrode
- 8 transparent substrate

【書類名】 図面

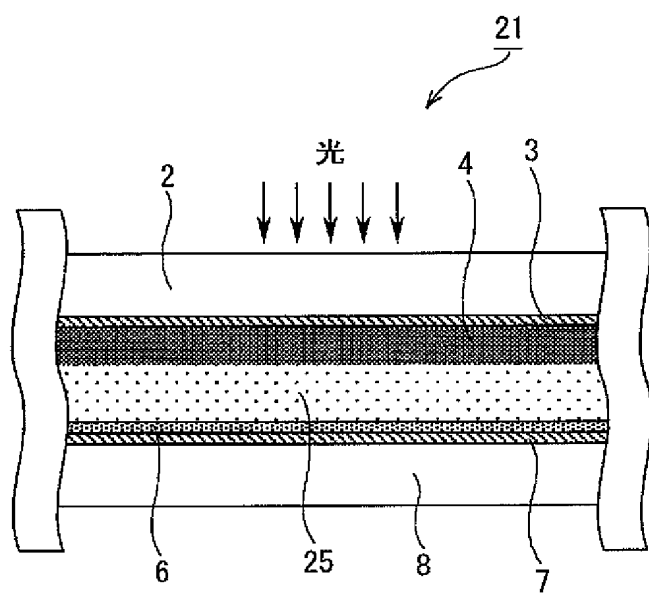
【図 1】



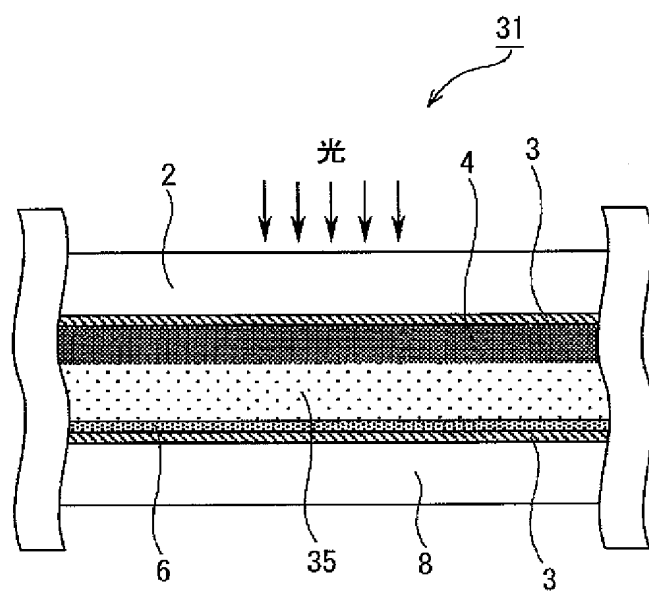
【図2】



【図3】



【図 4】



[Name of Document] DRAWINGS

[FIG. 1] LIGHT

[FIG. 2] LIGHT

[FIG. 3] LIGHT

[FIG. 4] LIGHT

[Name of Document]        ABSTRACT

[Abstract]

[Object]    To provide a solid electrolyte having superior conductive properties and reliability, a photovoltaic device using this electrolyte, and a method for manufacturing methods thereof.

[Solving Means]    A solid electrolyte of the present invention is a solid electrolyte having an electrolyte composition and a matrix polymer. The matrix polymer is formed by polymerization using a polyaddition reaction between a first compound having at least two isocyanate groups and a second compound having at least two nucleophilic groups containing active hydrogen, and the polymerization is performed after a precursor for the matrix polymer is brought into contact with a surface on which the solid electrolyte is to be formed.

[Selected Figure]        Fig. 1